

**Semiconductor Nanoplatelet Heterostructures Enhanced
via Combinations of Colloidal Atomic Layer Deposition
and Hot Injection Shell Growths**

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By
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SEMICONDUCTOR NANOPATELET HETEROSTRUCTURES ENHANCED
VIA COMBINATIONS OF COLLOIDAL ATOMIC LAYER DEPOSITION AND
HOT INJECTION SHELL GROWTHS

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July 2019

We certify that we have read this thesis and that in our opinion it is fully adequate, in scope and in quality, as a thesis for the degree of Master of Science.

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ABSTRACT

SEMICONDUCTOR NANOPATELET HETEROSTRUCTURES ENHANCED VIA COMBINATIONS OF COLLOIDAL ATOMIC LAYER DEPOSITION AND HOT INJECTION SHELL GROWTHS

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One of the most promising families of semiconductor nanocrystals in colloidal optoelectronics and nanophotonics is considered to be colloidal quantum wells, also commonly referred to as nanoplatelets (NPLs). Possessing an atomically flat structure, NPLs feature unique properties including spectrally-resolved and tunable light-hole- and heavy-hole transitions accompanied by their respective giant oscillator strengths. CdSe, CdS and CdTe, making the first colloidal NPLs synthesized in core-only structure, portray distinct qualities necessary for light-harvesting and -generating applications. However, going beyond the core structure, there are many properties that are highly enhanced by growing crown and/or shell layers around core NPLs. While the crown growth takes place anisotropically in lateral directions, the shell layer covers the entire NPL surface, combinations of which enable NPL heterostructures in new architectures. Depending on the electronic alignment of parts of the NPL heterostructure and the resulting confinement of electron-hole wave functions, these hetero-NPLs can be type-I or type-II. In type-I electron-hole pairs are confined in the core-NPL and recombination occurs in a direct pathway. In type-II electron-hole wave function is separated into different semiconductor layers, resulting in spatially indirect recombination. In this thesis, we synthesized and showed thin- and thick-shell grown heterostructures of type-I CdSe/ZnS NPLs using hot-injection (HI) for the first time particularly for these semiconductor NPLs. Unlike the typical colloidal atomic layer deposition (c-ALD) technique, which produces NPL heterostructures with low quantum yield (QY) and low chemical and optical stability, our approach yields CdSe/ZnS NPLs of almost unity (100%) quantum yield (QY) and improved chemical stability, tested by washing the same samples rigorously up to 6 times with ethanol with little change observed in the QY. Additionally, unparalleled thermal and optical aging endurance is achieved in aging tests. These tests experimentally demonstrated

that, elevated to 400 K, HI thick-shelled NPLs can retain up to 65% of their emission intensity in the colloidal form and 52% of that in the film. This level of high stability creates a great opportunity for employing these NPLs for high-temperature applications. Also, in the thesis, we synthesized and studied CdS/CdSe core/crown, CdS/CdZnS core/c-ALD shell-grown and CdS/CdSe/CdZnS core/crown/c-ALD shell-grown heterostructures of NPLs. Here the starting-template CdS NPLs are considered to be unique in terms of their emission in the blue region, which may open up new opportunities for NPL lasing in this spectral region. Nominally CdS NPLs are folded due to great lateral sizes. However, in this research work, when coated with crown and shell layer, these particles unfold. The unrolled CdS/CdSe core/crown NPLs are found to exhibit relatively higher QY up to 15-20% in its class of CdS core-seeded NPLs. The findings of this thesis reveal that such heterostructures of the NPLs are very rich in terms of variety of the quantum architectures one can achieve using them as working model systems.

Keywords: Colloidal quantum wells, nanoplatelets, core/crown, core/shell, hot-injection, colloidal atomic layer deposition.

ÖZET

KOLOİDAL ATOMİK KATMAN KAPLAMA VE SICAK ENJEKSİYON KABUK KAPLAMA YÖNTEMLERİNİN BİRLEŞİMİYLE GELİŞTİRİLEN YARI İLETKEN NANOLEVHA HETEROYAPILARI

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Nanolevhalar olarak da bilinen kolloidal kuantum kuyuları, kolloidal optoelektronik ve nanofotonik alanlarındaki önemli yarı iletken nanokristal malzeme gruplarından biridir. Atomik boyutları ve levhayı andıran düz yapıları sayesinde bu malzemelerin optik özellikleri istenildiği şekilde değiştirilebilir ve yine aynı sebepten ötürü bu malzemeler dev osilatör gücü gibi benzersiz özelliklere sahiptirler. CdS, CdSe ve CdTe, çekirdek şeklinde sentezlenmiş ilk yarıiletken nanolevhalar, ışık hasatı ve üretimi uygulamaları için gerekli özellikleri sergilerler. Ancak, bu optik ve kimyasal özellikler, çekirdeğin etrafında taç ve kabuk tabakaları büyütülerek iyileştirilebilir. Taç katmanı büyümesi yanal yönlerde gerçekleşirken, kabuk katmanı nanolevha yüzeyinin tamamını kaplar. Bu iki farklı katman büyütme tekniği birlikte kullanılarak yeni nanolevha heteroyapıları oluşturulabilir. Elektron ve deşik dalga fonksiyonlarının elektronik lokalizasyonuna bağlı olarak, bu hetero-nanolehvaları Tip-I ve Tip-II olarak ayrılır. Eksiton çiftlerinin sadece bir bölgeye hapsediği durumda Tip-I yapı oluşur ve rekombinasyon doğrudan gerçekleşir. Elektron ve deşik farklı yarıiletken katmanlarında lokalize olduğunda Tip-II yapı oluşur ve dolaylı rekombinasyon gerçekleşir. Bu tez çalışmasında, Tip-I yapıya sahip CdSe/ZnS ince ve kalın-kabuk nanolevha heteroyapıları ilk defa sıcak enjeksiyon yöntemi kullanılarak yapılmıştır. Düşük kuantum verime ve düşük kimyasal ve optik kararlılığa sahip heteroyapılarla sonuçlanan kolloidal atomik katman kaplama tekniğinin aksine, sıcak enjeksiyon yöntemi ile neredeyse %100 kuantum verime sahip ve en fazla 6 defa etanolle yıkadıktan sonra bile kuantum verimi neredeyse değişmeyen nanolehvalar üretilmiştir. Öte yandan sıcak enjeksiyon yöntemi ile sentezlenen kalın-kabuk nanolevhaların, 400 K sıcaklığında kolloidal şeklinde fotoluminesans verimliliğinin %65'ini ve film şeklinde fotoluminesans verimliliğinin %52'sini koruduğunu gösteren yaşlanma

testleriyle benzersiz termal ve optik yaşlanma dayanıklılığı tespit edilmiştir. Bu yüksek seviyedeki dayanıklılık, bahsi geçen nanolevhaları yüksek sıcaklık uygulamaları için kullanılmasına fırsat yaratır. Tez çalışmasının bir başka kısmı, CdS nanolevhalarının ve onların CdS / CdSe çekirdek / taç, CdS / CdZnS çekirdek / c-ALD kabuk ve CdS / CdSe / CdZnS çekirdek / taç / c-ALD kabuk heteroyapılarının sentezini de kapsamaktadır. Nanolevhalar arasında CdS yalnız mavi bölgede ışıma yapmaktadır, bu özellik lazer teknolojisinde yeni kapılar açabilir. Normalde, büyük yanal boyutları nedeniyle CdS yapısal anlamda katlanmaktadır. Bu araştırma çalışmasında, taç ve kabuk tabakası ile kaplandığında, bahsedilen nanolevha katlanmasının önüne geçilmektedir. Bu CdS / CdSe çekirdek / taç nanolevhalarında, yalnızca çekirdek yapıdaki CdS nanolevhalarıyla karşılaştırıldığında, kuantum verimi %15-20'ye kadar artış göstermektedir. Nihai bulgular, bu nanolevha heteroyapılarının çeşitli kuantum mimarileri yönünden oldukça zengin olduğunu göstermektedir.

Anahtar sözcükler: Koloidal kuantum kuyuları, nanolevhalar, çekirdek/taç, çekirdek/kabuk, sıcak enjeksiyon, koloidal atomik katman kaplama.

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Chapter 1

Introduction

Starting from the 80s of the last century, semiconductor nanocrystals (NCs) have become one of the most extensively studied areas of the nanoparticle realm. One of the major breakthroughs came with the discovery of quantum size effect in NCs [3]. Semiconductor NCs, specifically, colloidal quantum dots (QDs) whose radii are smaller than or comparable to the bulk exciton Bohr radius, constitute a class of materials intermediate between molecular and bulk forms of matter. A straightforward and easy colloidal pathway to obtain monodisperse NCs was introduced by Bawendi [4]. The advancement in the synthesis of various shapes of semiconductor nanoparticles continued with nanorods, nanowires, nanotetrapods and finally leading to nanoplatelets (NPLs).

Nanoplatelets (NPLs), the so-called colloidal quantum wells, are free-standing quasi-two-dimensional nanosheets having strong one-dimensional quantum confinement. They feature a giant oscillator strength transition. They are fast colloidal fluorescent emitters exhibiting nanosecond lifetime at temperatures approaching cryogenic [5][6]. One of the advantages of NPLs over QDs is their large volume relative to QDs. Thus, NPLs show weaker Coulomb blockade effects and hence, permits enhanced charge injections without affecting quantum confinement. While synthesizing QDs requires precise control over 3 directions, in the case of NPLs only vertical thickness should be precisely controlled and manipulated. As a result of uniform vertical thickness, NPLs have narrow photoluminescence (PL), which is highly favorable for applications such as LEDs that need high color purity.

Nowadays, the main synthesis pathways for 2D materials include molecular beam epitaxy, exfoliation, and colloidal synthesis in solution [7][8]. Unlike the 2D systems which are produced through gas phase deposition, the number of atomic layers of nanoplatelets can be tuned down to single atomic layers, which is why they are often referred to as atomically flat [5].

In recent years, new NPL heterostructures (core/crown and core/shell) have been made to enhance chemical and optical properties of the core NPL structures such as photostability and PL QY. In most cases, especially those involving high temperature growth, shell layer leads to a strong enhancement of the PL QY [9]. On the other hand, in the core/crown structure outer crown layer acts as a exciton funnel that gathers excitons and channels them to the core [10], which can be explained by the difference in dielectric constant of the core and crown materials.

In this thesis work, colloidal CdSe and CdS NPLs were synthesized and covered by using combinations of hot-injection (HI) grown ZnS shell layers and CdSe crown, colloidal atomic layer deposition (c-ALD) grown CdZnS layers. The main motivation of the synthesis of CdSe/ZnS core/HI shell is to produce environmentally and thermally stable shell structures. The HI shell deposition was mostly implemented by using CdS as an outer shell layer. However, because of the substantial distinction in CdSe and ZnS band gaps, this shell can produce better passivation. Therefore, enhanced NPLs can be synthesized for applications requiring elevated temperatures. When it comes to CdS, the motivation stems from the fact there is not much application of NPLs emitting at blue region of the colour spectrum. New core/crown recipe might improve QY and c-ALD shell growth and assist in uniform passivation of the rolled CdS NPLs and thus, opening up for the further improvement.



Figure 1.1: Various NPLs synthesized at our laboratory emitting at different wavelengths of the color spectrum.

Chapter 2

Synthesis methods of colloidal nanoplatelets

2.1 Core synthesis

Efficient synthesis of colloidal core NPLs is a critical step for heterostructures to be grown on them. In recent years, among most extensively studied NPLs, CdSe, CdTe and CdS cores have been important in terms of their potential applications. They have tunable and narrow bandgap emission thanks to their monodispersity. To obtain pure 2D NPL cores with enhanced optical features and chemical stability, a good control of the reaction parameters (e.g., capping agents, molar ratio, temperature, time), along with the right choice of appropriate precursors (e.g., transition metals and chalcogens), is crucial. As a result, the intrinsically desired anisotropic growth can be preserved. Usually, -chloride , -carbonyl and -carboxylate precursors are employed as high-yielding metal precursors [5], [8], [11], [12]. The most distinguished and thoroughly explored core NPLs are CdSe nanoplatelets. Their thickness ranges from 2 to 8 monolayers (MLs) [13]. Among them 4 ML CdSe synthesis is the simplest, the least time-consuming, and requires the least material input. The as-synthesized NPLs are atomically flat, they have precise vertical thickness, and require minor cleaning by size-selective precipitation through the usage of polar solvents such as ethanol and methanol [5], [11], [12].

The synthesis of NPLs, in general, is carried out in an apparatus called Schlenk line consisting of two manifolds with numerous ports (Figure 2.2). One manifold is usually connected to vacuum pump and the other one to neutral gas (typically nitrogen or Argon). 4 ML CdSe core NPL synthesis usually involves Cd salt with a long hydrophobic chain, a non-coordinating solvent and a Se source in the form of elemental Se. The major CdSe synthesis pathway starts with preparing $\text{Cd}(\text{myristate})_2$ precursor. It is made by mixing Cd nitrate tetrahydrate and sodium myristate solution

in methanol at room temperature and then filtering the residual Cd salt after the synthesis. The next stage involves a quite well-defined sequence of steps with minor variations:

- Adding 340 mg Cd(myristate)₂, 24 mg Se and 30 mL octadecene (ODE) to a 100 mL flask,
- Degassing for some time,
- Adding Cd precursor with a relatively shorter chain at an elevated temperature at which the solution acquires a golden color, indicating the nucleation of particles having 4 ML thickness (usually around 195 °C),
- Starting a timer around 220 °C when Cd(myristate)₂ starts to decompose and the actual growth starts,
- Continuing the reaction depending on the targeted lateral size (i.e., keeping longer for large area NPLs), and
- Finally, quenching the reaction with oleic acid (OA).

The synthesis of CdS NPLs was first discovered by Dubertret's group [5]. In their case, unlike CdSe NPLs, CdS NPLs did not have sharp features and thus, had different subpopulations. They used Cd myristate as the Cd precursor and Zn acetate as the capping agent. Later, Peng et al. suggested different synthesis routes using Cd acetate dihydrate and, instead of degassing, they applied Argon bubbling [14]. Also, they used oleic acid as the ligand but also worked with other ligands such as myristic acid, stearic acid and melissic acid.

The major reaction stages in the NPL synthesis can be divided into the conversion of precursors to seeds, the nucleation and the growth. Through investigations using optical methods and transmission electron microscopy (TEM) it is suggested that the initial seed formation dictates thickness of NPLs with lateral extensions later in the subsequent growth stage. There are few hypothesized growth mechanisms for NPL core formation as reported by Dubertret et al. According to one potential mechanism, the lateral extension happens when the seeds self-organize on the seed level or in patches. In another proposed mechanism, the seeds extend in lateral direction through the continuous reaction of precursors (Figure 2.1) [15].

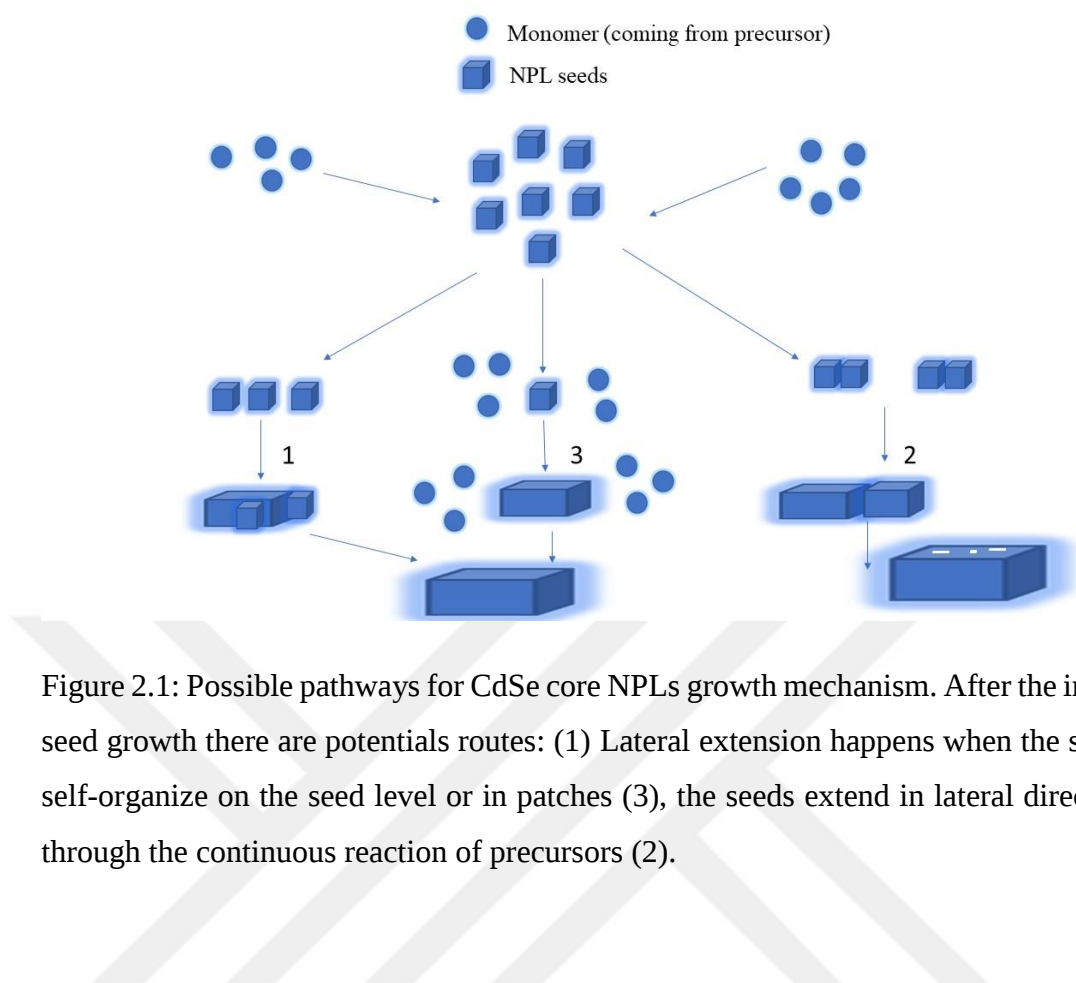


Figure 2.1: Possible pathways for CdSe core NPLs growth mechanism. After the initial seed growth there are potentials routes: (1) Lateral extension happens when the seeds self-organize on the seed level or in patches (3), the seeds extend in lateral direction through the continuous reaction of precursors (2).



Figure 2.2: The Schlenk line apparatus and utilities employed for the synthesis under non-ambient conditions at our laboratory.

2.2 Crown growth

To improve optical properties of NPLs, anisotropic core/crown NPL heterostructures can be synthesized. The crown layer is an extension in the lateral direction and is atomically flat with the same thickness as the original core NPL. It is usually grown by choosing shorter-chain carboxylic acids, using different temperatures (usually lower) and making precursors more concentrated compared to those of the core. Depending on the size of the core, the resulting core/crown heterostructures can display properties related to either the core semiconductor or the crown semiconductor. According to the band structure of the formed heteronanoplatelets, the core/crown NPLs are divided into 2 types: type-I and type-II (Figure 2.3). In the type-I NPLs, the core is covered with a wider band gap material and both charge carriers (hole and electron) are confined inside the core NPL [10], [16]. The crown coating acts as an optical antenna for excitation funnelling the photogenerated NPLs and increasing the NPL's emission. CdSe/CdS NPLs are a common example of type-I heteronanoplatelets. In the type-II nanoplatelets, electron and hole are confined in different parts of the heteronanoplatelets. As a result, charge carriers are separated between the core and crown regions [16], [17]. CdSe/CdTe heteronanoplatelets are a good example of type-II core/crown NPLs [17].

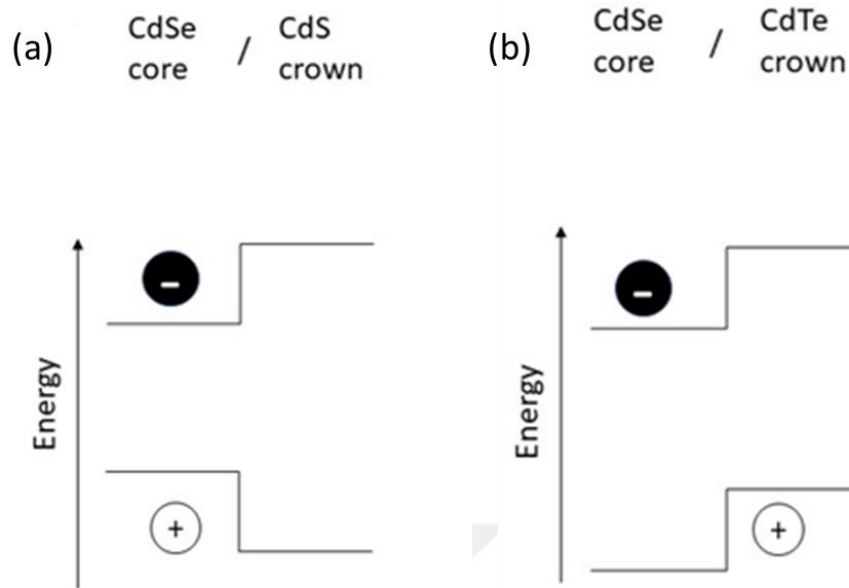


Figure 2.3: Schematic depiction of processes going on upon excitation of (a) type – I CdSe/CdS and (b) type – II CdSe/CdTe core/crown heterostructures.

The main synthesis process starts with the core synthesis as explained in Section 2.1 and its purification is carried out carefully beforehand. Then, in the case of CdS coating, a Cd crown precursor is synthesized separately. Usually, a mixture of Cd(acetate)₂ dihydrate, OA and ODE are heated up to 130 °C under constant stirring with a magnet bullet in a round bottom flask. Then, a cycle of stirring and sonication is initiated, which provides complete dissolution of OA and Cd acetate salt. Then, at a certain optical density (O.D.) of core NPLs, ODE and OA depending on the desired final lateral size of the core/crown NPLs, a certain amount of Cd crown precursor is degassed for 20-30 min until hexane (or any other core solvent) evaporates. After changing to gas, at high temperature values, chalcogen precursors are added either instantly or via gradual addition through a syringe pump. The final product is precipitated by the addition of ethanol (or other solvents) in a centrifuge machine.

2.3 c-ALD shell growth

While the crown layer is a lateral extension to the core NPLs, a shell fully covers the NPL surface, mainly those lateral (top and bottom) parts that are not protected by the crown layer. The shell plays the role of a barrier between the core and the environment, thus making the NPLs less sensitive to changes in the surrounding, surface chemistry, and diminishing possibility of photo-oxidation. Also, the shell growth allows for better passivation of the surface trap states, which in turn enhances fluorescence quantum yield (QY).

The first reported core/shell NPL synthesis was carried out using colloidal atomic layer deposition (c-ALD). Atomic layer deposition is used widely for the deposition of thin films made from metals, semiconductors, or even dielectric oxides on various substrates. The key property of ALD is that it enables to deposit the exact number of layers on substrates under ambient temperature conditions. Surface reactions are limited by a single layer and before starting a new layer the system is purified to get rid of the precursors from the previous step [18]. In c-ALD, the same set of principles are applied to colloidal nanocrystals syntheses, enabling the precise control over number of layers. For each semi-layer, excess of reactants is used which is later removed by a purification process, e.g., precipitation with polar solvents.

The principal idea of c-ALD is based on the exchange of NPLs and precursors between polar and non-miscible non-polar solvents [19]. It is, primarily, divided into 3 subgroups: c-ALD by phase transfer of nanoparticles, c-ALD in nonpolar phase without phase transfer, and c-ALD growth in polar phase. Generally, chemicals such as formamide (FA) and N-methyl formamide (NMF) are used as a polar medium and toluene and hexane are used as a non-polar solvent. The typical c-ALD route employed for NPLs is the polar phase transfer, which can be summarized in the following steps:

- First, colloidal NPL cores are synthesized,
- Then they are cleaned at least 2 times to remove most of the surface capping agents and dissolved in non-polar solvents,
- Later an anionic semi-layer precursor (such as ammonium sulphide for S^{2-} layer) and a polar solvent (e.g., NMF) are added and the resulting mixture is agitated in a vortex mixer for few minutes (here the color change indicates the transition from one solvent to another and capping with the new semi-layer),

- Non-polar solvent is carefully decanted and the solution is cleaned from excess precursors by centrifugation with acetonitrile and toluene,
- The precipitate is dissolved in the polar solvent, cationic precursor such as Cd acetate dihydrate solution in NMF is added and vortexed for few minutes,
- The unreacted precursor is again removed by centrifugation after the addition of toluene,
- For each extra complete layer, the above steps are repeated,
- The resulting core/c-ALD shell NPLs are coated with ligand layer (OA) by the addition of some of ligands and non-polar solvents and slightly heating,
- Finally, after some time is given for the final solution to stabilize, it is size-selectively cleaned by ethanol or other polar precipitating solvents.

In Figure 2.4, the products of a typical core/ multi-layer c-ALD shell synthesis are shown. It can be seen that the addition of each monolayer in the shell changes the color of the NPL solution, resulting in a red-shift with each addition.



Figure 2.4: A photo image of our CdSe/CdS core/ c-ALD shell heteronanoplatelet samples coated with 1, 2, 3, 4, 5- and 6-ML c-ALD shells at our laboratory: from left to right.

2.4. Hot–injection shell growth

Although c-ALD shell growth is considered to be precise in terms of controlling the number of grown shell layers, it takes a lot of time and resources. Typically, the synthesized NPLs have low QY and are not much chemically stable. One of the possible solutions is the implementation of shell synthesis at a relatively high

temperature. Recently, heat-up method has been suggested for the semiconductor shell growth around NPLs [20]. This step was carried out at temperatures ranging from 100-150 °C. However, a new approach yielding semiconductor NCs at a high luminescence QY and with monodispersity in size and shape was needed. For this reason, hot-injection method is a promising candidate. The hot-injection (HI) was first introduced in 1993 in the synthesis of CdSe, CdS and CdTe quantum dots. Later, it was used for the growing shell around colloidal CdSe QDs [21]. Depending on the shell precursors, the used temperatures can range from 150 to 350 °C. In the case of colloidal NPLs, we developed and showed the hot-injection reaction pathway for CdSe/ZnS core/shell NPLs [2], which will be discussed in the following chapters of this thesis. This approach uses octanethiol as S source, which is more active around 300 °C and Zn acetate as Zn metal source.

Chapter 3

Characterization methods of colloidal nanoplatelets

3.1. Optical characterization

Typically, the first step in characterization of the as-synthesized colloidal NPLs is determination of their optical properties. These properties can be manipulated and tuned for the potential applications of NPLs in photovoltaic, lasing and display devices. By means of optical characterization one can infer their electronic structures, predict their shape, evaluate their monodispersity level, understand their surface chemistry and their crystal structures [15]. It is obvious from these analyses that NPLs exhibit superior qualities such as giant oscillator strength, narrow emission spectrum, reduced Auger recombination and inhomogeneous broadening [5].

3.1.1. UV-Vis spectroscopy

One of the key factors defining NPLs and colloidal semiconductor nanocrystals, in general, is their absorption features. The absorption spectrum tells a lot about their purity, shape, size, and possible area of applications [22]. Therefore, UV-Vis spectroscopy, which utilizes electromagnetic radiation ranging from 200 to 1100 nm, can be considered one of the primary tools for characterizing and hence understanding NPLs. The reason for choosing this wide range is that: while infrared radiation causes only vibrational transitions, UV and visible electromagnetic spectrum with higher photon energy causes electron transitions. For example, a molecule is exposed to light, some of electrons in the ground state in the Highest Occupied Molecular Orbital (HOMO) jump to the Lowest Unoccupied Molecular Orbital (LUMO). The gap between these HOMO and LUMO states defines which wavelength of light is absorbed by the electrons of the matter of interest. The general equation governing the absorption level is given by Beer-Lambert Law:

$$A = \log_{10} \frac{I_0}{I} = \epsilon * l * c \quad (3.1)$$

where the symbol A shows the absorbance of the sample, ϵ is a proportionality constant, which is called molar absorptivity or molar extinction coefficient, l is the length of the light path (optical path inside the cuvette), c is the concentration of the sample, and I_0 and I denote the intensity of the light passing through the sample cuvette and the reference cuvette, respectively.

In the experiments, Cary 100 was used as a UV-Vis spectrometer. The general scheme of a UV-Vis spectrometer is shown in Figure 3.1. The working principle can be summarized as follows:

- At first, the light from the lamps (tungsten and deuterium arc lamps) passes through the filter reaching the monochromator where it is separated into different wavelengths,
- Then, the light is divided into two passes by the beam splitter: a part of it goes through the sample cuvette and a part goes through the reference solution usually consisting of pure hexane or toluene,

- Finally, the beam passes through the detectors (which are photodiodes in the figure) and the final output showing absorbance vs wavelength (nm) is recorded.

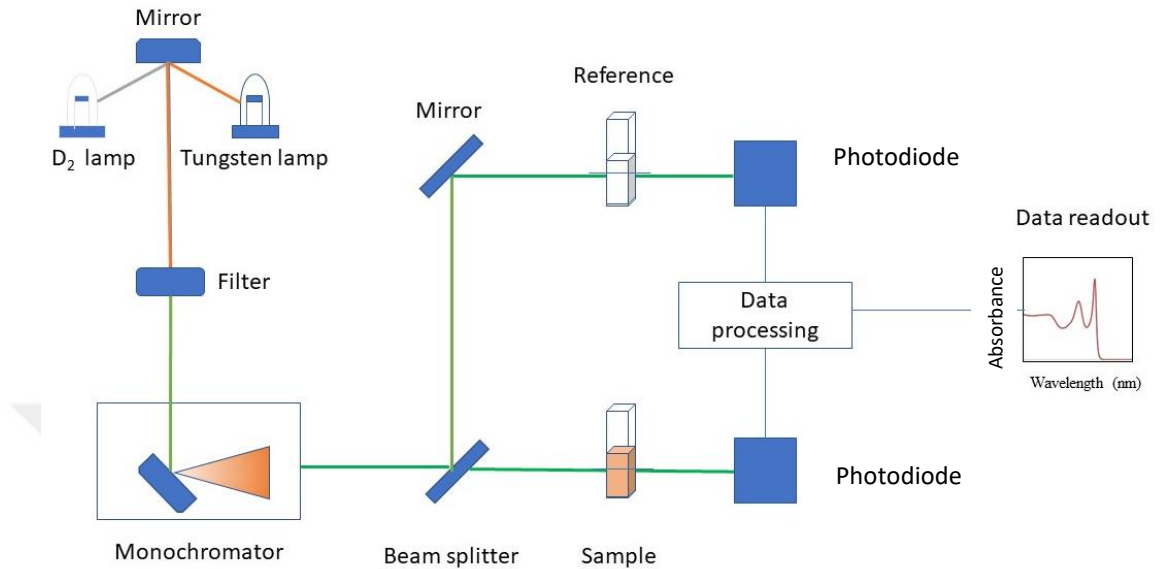


Figure 3.1: General scheme for a typical double-beam UV-Vis spectrometer.

The typical absorption spectrum of 4 ML CdSe NPLs is depicted in Figure 3.2. There are certain electron transitions clearly visible in the spectrum: the electron/heavy-hole transition at around 513 nm (very slight deviation depending on the lateral size) having the lowest photon energy and the electron/light-hole transition at around 480 nm having the highest photon energy value [5].

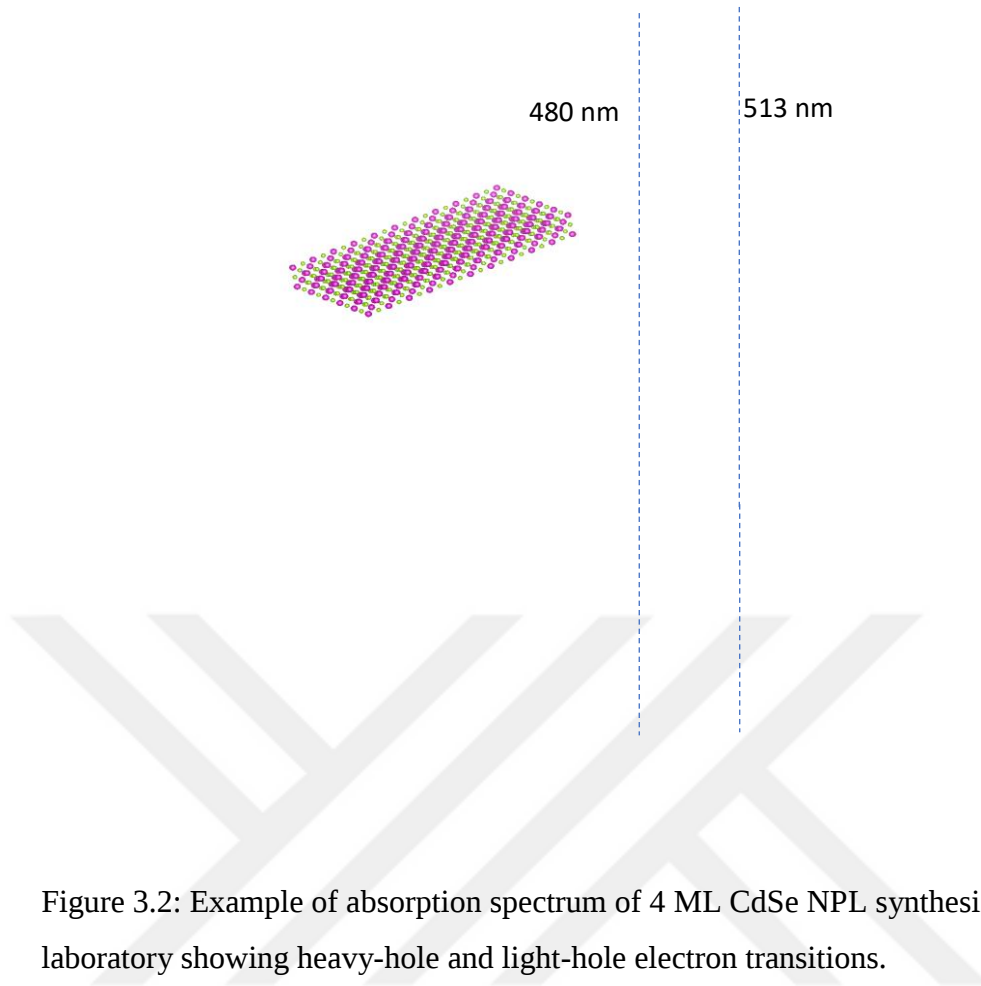


Figure 3.2: Example of absorption spectrum of 4 ML CdSe NPL synthesized at our laboratory showing heavy-hole and light-hole electron transitions.

3.1.2. Fluorescence spectroscopy

Another crucial instrument for analysing optical properties of NPLs is the fluorescence spectrophotometer. Fluorescence, by its nature, is a type of the more general concept of photoluminescence. In fluorescence, the singlet excitation occurs where electron preserves its initial spin and it decays faster compared to phosphorescence where emission occurs between any two energy levels that differ in their respective spin states. In this research work as a fluorescence spectrophotometer Cary Eclipse model of Varian company was used. Its main parts consist of a source (typically Xe arc lamp), a sample compartment and a monochromator system which defines excitation and emission wavelengths. Such an instrument, which employs a monochromator system instead of simply filters, is referred to as a spectrofluorometer.

A typical PL spectrum of 4 ML CdSe NPLs has a sharp peak at around 513 nm with 6-10 nm full-width at half-maximum (FWHM) at room temperature. This narrow FWHM is unique for NPLs which implies high monodispersity of the synthesized NPLs in their vertical thickness [5], [15], [20].



Figure 3.3: Example of PL emission spectrum of 4 ML CdSe NPL sample synthesized at our laboratory.

3.1.3. Quantum yield measurement

One of the most important factors defining colloidal nanocrystals' applications is their fluorescence quantum yield (QY). This is defined as the number of emitted photons over the number of absorbed ones. QY is also often referred to as quantum efficiency because it defines the efficiency of conversion of absorbed light to emitted light. QY can be determined in two ways: by absolute measurements using an integrating sphere setup or by relative measurements using the fluorescent standard of a known reference dye [23], [24]. When conducting relative QY measurements, the first step is the choice of an appropriate dye with a similar emission range. Then, the absorbance of the

reference dye and that of the sample under test are adjusted so that they are similar and below 0.1 absorbance value. As the next step, PL emission is measured and the integrated PL emission is calculated (by integrating the area under the curve). Later, by using the calculated integrated PL emissions of the NPL sample and that of the reference dye based on the already known QY value of the reference dye, the QY of the NPL sample is computed.

To perform absolute QY measurements, the dilute sample is loaded in special cuvettes and placed inside the integrating sphere device. Depending on the emission of the sample, the monochromator is used to adjust excitation created by the light source. Four different spectra are collected in four different configurations: (1) the blank spectrum when the integrating sphere is empty and the light source is open, (2) the sample spectrum when the sample is placed inside the sphere and is excited directly by the light source, (3) the sample is still inside but indirectly excited by the light source, and finally (4) the sample is inside, however, there is no light coming inside. From these four spectra based on the method developed by Mello et al. the absolute QY values are determined [15]. Special attention should be paid for using reasonable concentration values.

3.2. Structural characterization

To analyse shape, size, structure, elemental composition and crystalline structure, various instruments including transmission electron microscope (TEM), x-ray diffractometer (XRD) and x-ray photoelectron spectrometer (XPS) are being used. Obtaining information about the shape and size sheds light on the possible growth of crown and shell layers, while the elemental structure helps to understand more about the growth mechanism and if reactions go each time as planned.

3.2.1. Transmission Electron Microscopy (TEM)

One of key discoveries for the modern nanoscience was the invention of TEM in 1931 by Ernst Ruska at the University of Berlin, along with Max Knoll. The advantage of electron microscopes over light microscopes is the small de Broglie wavelength of electrons as opposed to light wavelength. This allows images with much higher

resolution. Unlike the scanning electron microscopy (SEM) which, primarily, produces images through secondary electrons and electrons back scattered from the surface of the sample, in TEM the images are obtained via information directly collected through the transmitted electrons. Therefore, TEM samples should be thin (typically less than 100 nm). The resolution of TEM is much higher than SEM and can reach 0.1 nm in some modern instruments. Figure 3.4 shows the essential parts of a TEM instrument commonly used in the industry or academia. TEM instruments typically include: a column where the major process is adjusted, left-hand (LH) and right-hand (RH) control panels, binoculars, viewing fluorescent screen, main display and hardware items such as utilities, vacuum pumps, and electronics (Figure 3.4). The column part can further be divided in few crucial parts: an electron gun, condenser lens, a sample stage, objective lenses, objective lens aperture, intermediate lens, project lens, and fluorescent screen (Figure 3.5). Another important feature of conventional TEM hardware is an energy-dispersive X-ray spectrometer (EDS). This is employed for elemental analysis and can be used in dot, line or mapping modes. TEM is constantly cooled with liquid nitrogen while operating and cryo-cycle is turned on whenever it is not used. There are typically three accelerating voltages used: 100 kV, 200 kV and 300 kV.

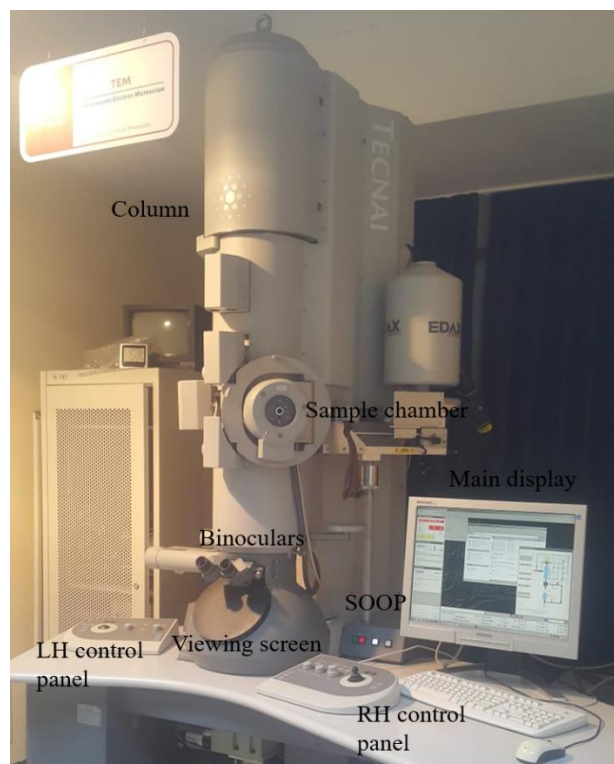


Figure 3.4: A typical TEM.

For the structural characterization carried out in this thesis work FEI Tecnai G2 F20 TEM was used. The NPL samples were prepared by first cleaning the liquid sample with ethanol and acetone solvents several times to get rid of the most ligands, which would otherwise burn under the electron beam and make it impossible to focus image for higher magnification and better resolution. Commonly 40-50 μL of the very dilute final solution was dropped on the TEM grid CF300-CU-50 and dried thoroughly. Usually, high-angle annular dark field scanning transmission electron microscopy (HAADF-STEM) mode was used to image NPLs to understand their shape and measure size. The accelerating voltage was usually set to 200 kV. Some typical TEM images of the CdSe and CdS NPLs are depicted in Figure 3.6.

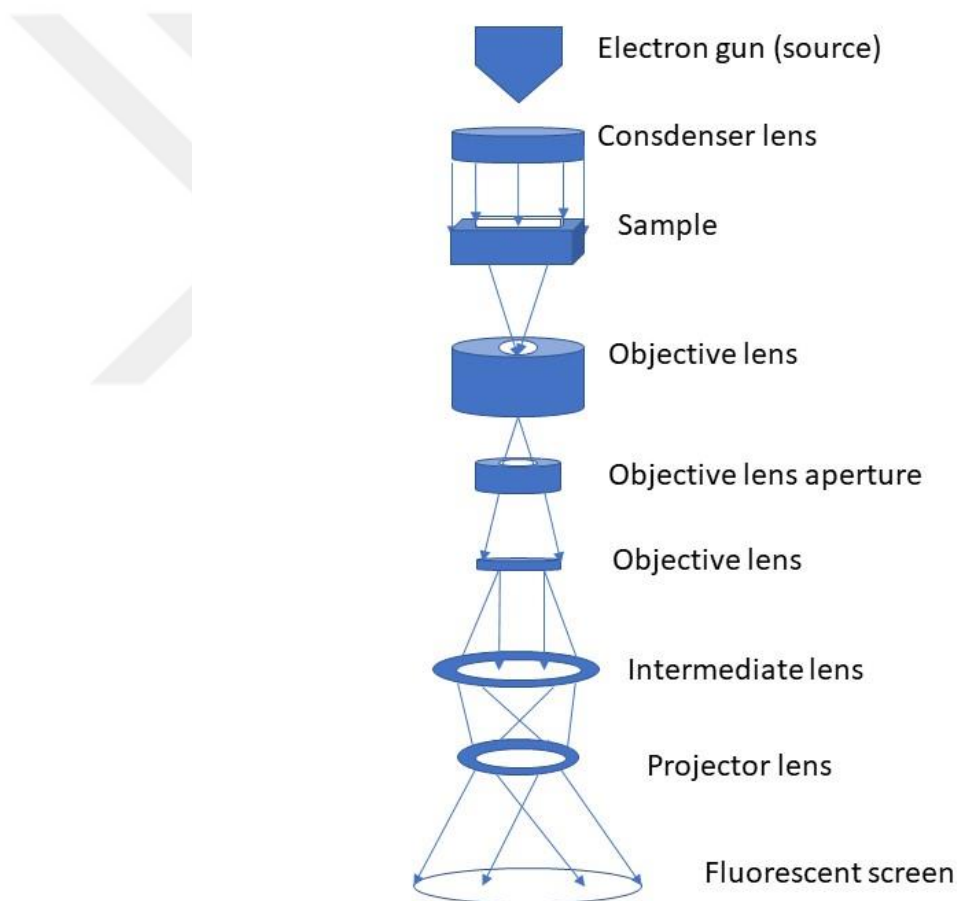


Figure 3.5: Schematic representation of TEM column.

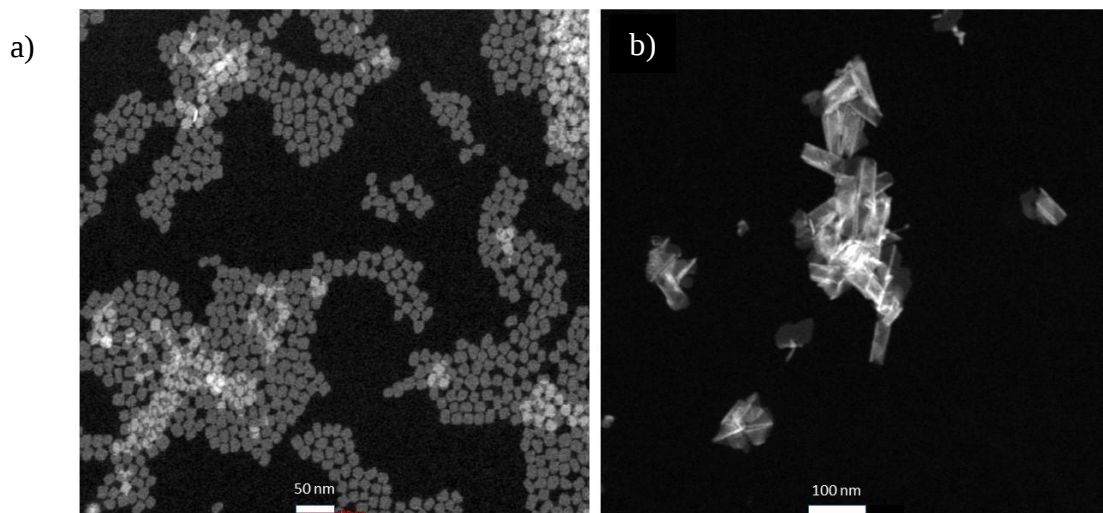


Figure 3.6: Examples of HAADF-STEM images of (a) CdSe NPLs and (b) CdS NPLs synthesized at our laboratory.

3.2.2. X-ray powder diffraction spectroscopy

X-ray diffractometer (XRD) is an essential instrument for analysing crystalline matter, and hence, it is also an important tool for characterization of colloidal semiconductor nanocrystals, especially, NPLs. Though there are numerous types of XRD, the primary one for NPL analyses is an X-ray powder diffractometer often used for polycrystalline sample analyses. The idea to use X-rays for analysing crystal structure was generated when Max von Laue discovered in 1912 that substances having crystalline structure rather than amorphous behave as three-dimensional diffraction gratings for X-ray wavelengths because of the equal spacing of planes in a crystal lattice. X-rays are generated by cathode ray tube. Later they are filtered to produce collimated monochromatic radiation and are directed toward the sample. The main working principle of XRD is based on the constructive interference of the X-rays incident onto the crystalline sample. This happens when the Bragg's condition is satisfied

$$n\lambda = 2d \sin \theta \quad (3.2)$$

where n is an integer number, λ is the wavelength, d is the spacing between two planes, and θ is the produced diffraction angle (Figure 3.7).

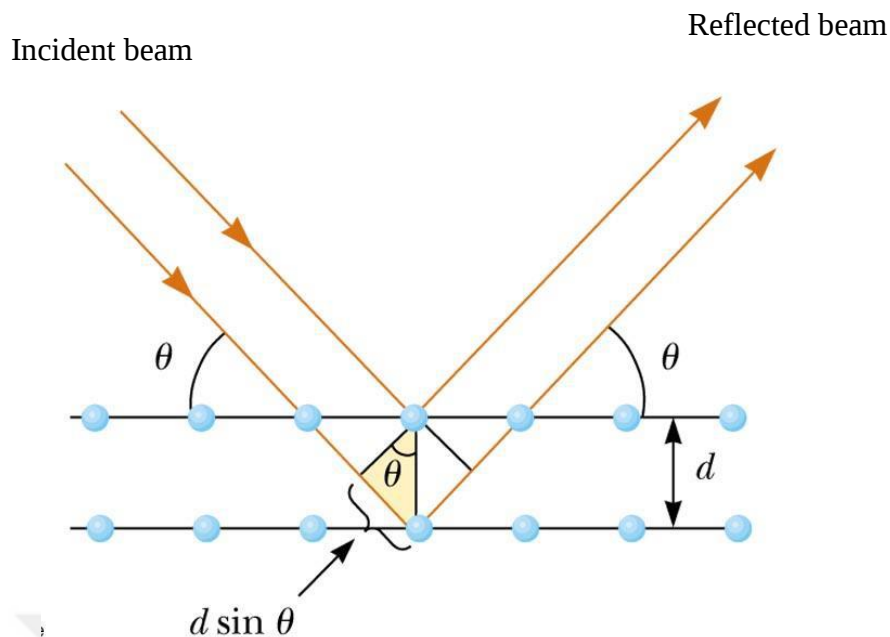


Figure 3.7: A schematic of Bragg's law working principle.

In our studies, XRD patterns of NPLs were used to prove that indeed the synthesized CdSe/ZnS have the corresponding crystalline structure. The typical zinc-blende (ZB) CdSe NPLs have well-defined XRD spectrum peaks for their (111), (220), (211) and (400) planes (Figure 3.8). Samples were prepared by drying NPLs and dropping the concentrated solution on the corresponding holders.

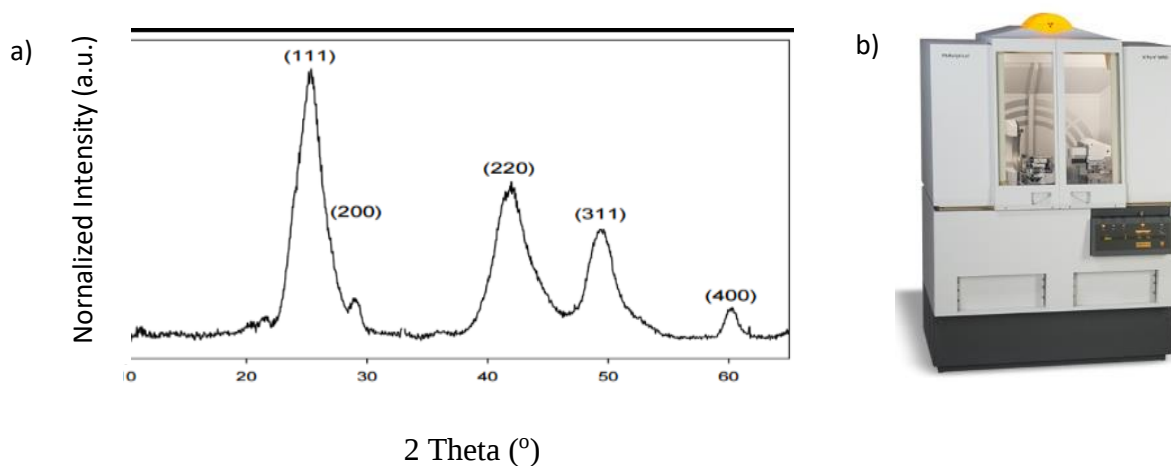


Figure 3.8: a) XRD spectra of CdSe NPLs, b) XRD instrument.

3.2.3. X-ray photoelectron spectroscopy

X-ray photoelectron spectrometer (XPS) is one of the useful tools to determine elemental structure and distribution of NPLs. It is a surface sensitive technique meaning it provides information only about the surface structure and to go deeper, one should etch the surface by means of ion gun. The major principle behind the working of XPS is photoelectric effect, which was first proposed by Hertz and later explained by Einstein as a way by means of which light manifests its particle nature. When light is shone on a sample, a photon can be absorbed by an electron and it can escape from the surface with a maximum kinetic energy of

$$E_k = h\nu - E_B - \Phi \quad (3.3)$$

where ν denotes the photon frequency, E_B is electron binding energy and Φ is a work function, which gives the minimum energy required to remove a delocalised electron from the surface.

The general process has few essential steps. X-ray having energy $h\nu$ is irradiated upon the sample surface to knock out surface electrons. Photons that have more $h\nu$ energy go deeper on the sample surface. Electrons which come from deeper levels are called photoelectrons, they are bound to the core level nucleus with binding energy E_B and the whole phenomenon is referred to as the photoelectric effect. Those generated free electrons possess the kinetic energy E_k after leaving the surface. It is this E_k that is detected and from that the binding energy E_B is calculated.

XPS results are usually used in conjugation with EDS spectra. As the XPS measurements are conducted under ultrahigh vacuum, the samples are dried and placed on special holders very carefully. Any possible contamination of the chamber should be avoided. The obtained spectra are shifted with respect to C 1s binding energy, which has a peak at 284.8 eV.

3.3. Aging tests

Though not as widely used as the other characterization methods aging tests are beneficial for understanding potential applications of colloidal nanocrystals including NPLs. They are often used for material quality testing in other fields, particularly industry to understand chemical and mechanical endurance of various materials. NPLs have high surface-to-volume ratio, which leads to rather reduced thermodynamic stability and degradation of optical properties [25][26]. Therefore, in applications which require elevated temperature conditions such as LEDs and lasers understanding thermal stability properties of NPLs is crucial, which can be conducted using various thermal and optical tests.

It is known that PL quenching for QDs happens at elevated temperatures and this leads to decline in optoelectronic features [27]. However, elevated temperature optoelectronic performance of NPLs is one of the critical aspects yet to be studied. To date there has not been much research carried out about it. Probably one of the most significant works was reported by Rowland et al [28], in which CdSe NPLs and CdSe/CdS core/c-ALD shell heteronanoplatelets were examined via static and transient spectroscopies with temperatures ranging from 300 to 700 K under vacuum. Rowland et al. also used in situ TEM which has a feature of temperature control and allows to look at the shape deformations due to temperature elevation. The authors reported that the PL Intensity decreased by about 8-10 times when going from 300 to 400 K, and concluded that it does not depend on the thickness of NPLs or whether there was CdS shell layer. However, the PL Intensity was fully recovered for this range when ambient temperature conditions were restored. Also, unlike QDs, NPLs showed detectable PL Intensity up to higher temperature values. They made a final conclusion that the main culprit behind the PL quenching upon temperature elevation is the PL quenching. Another crucial deduction thanks to TEM images was that morphology remained intact for 4 and 5 ML CdSe up to 560 and 630 K, nevertheless above those values that was deformation cannot be recovered. This was found, in turn, to contribute to the final irreversibility of PL emission. In this thesis work, new approaches were used encompassing both liquid and film thermal tests, as will be presented in the following chapters.

4. Shell growths of semiconductor nanoplatelet heterostructures

4.1. Chemicals

Chemicals used for the synthesis of CdSe NPLs include the following: 1-octadecene (ODE, 90,00 %), selenium (Se, 99,00 %), 1-octanethiol ($\geq 98,50$ %), zinc acetate (99.99 %), n-hexane (97,00 %), toluene (≥ 99.90 %), oleylamine (70 %), zinc acetate dihydrate (98 %), ethanol (absolute), oleic acid (OA, 90,00 %), Cd acetate dihydrate (98,00 %), sodium myristate ($\geq 99,00$ %), Cd nitrate tetrahydrate ($\geq 99,00$ %), Cd acetate (99.99 %), methanol (99.70 %), Cd acetate monohydrate (99,00 %), N-Methylformamide (NMF, 99.00 %), ammonium sulfide solution ($(\text{NH}_4)_2\text{S}$, 40-48 wt. % in H_2O), acetonitrile (≥ 99.90 %), sulphur (S, 99,00 %), and acetone (absolute). All of the substances were acquired from Sigma- Aldrich. No additional purification was applied.

4.2. CdSe core synthesis

The basic steps of the synthesis follow the procedure summarized in Chapter 2.1. However, the size-selective precipitation part is worth mentioning here. First of all, after adding 5 mL n-hexane for better dissolution, centrifugation at 6,000 rpm for 6 min

is carried out. The first precipitate mostly consists of unreacted Se and other compounds. Then, the remaining supernatant is divided into 2 parts. After adding 5-6 mL ethanol until the solution became blurry it is centrifuged at 6000 rpm for 6 min. Mostly 4 ML CdSe NPLs precipitate using these particular conditions. However, there still might be a little amount of 3 and 5 ML NPLs, QDs (requires checking with UV-vis and PL spectroscopy) and some OA. Therefore, the sample is kept for one day at the refrigerator. Later, the excess OA, some QDs and 5 ML NPLs are precipitated with 10,000 rpm centrifugation for 3-4 minutes. The precipitate is then disposed of. The hardest part is cleaning 3 ML NPLs, if any present. Some amount of ethanol is added until the liquid becomes very little blurry and it is precipitated at 6,000 rpm. Both the precipitate and the supernatant are checked for the presence of 3 MLs. If necessary, this latter cleaning step is repeated. Nonetheless, with each step more ligands on the surface are detached, and as a result, QY drops and concentration decreases. So, the better reaction goes, the fewer 3 ML NPLs are produced and the higher final quantum and chemical yields are. At the very end, typically, an additional 20 min 10,000 rpm centrifugation is conducted to prevent any non-fluorescent nanoparticles to remain in the final product.

The major modifications are observed in the CdSe NPL depending on the type and amount Cd precursor the shape changes. The shape of the as-synthesized NPLs turned out to be square, rectangular or irregular. This is presumed to be related to the amount water molecules present in Cd acetate, Cd acetate monohydrate and Cd acetate dihydrate precursors [11].

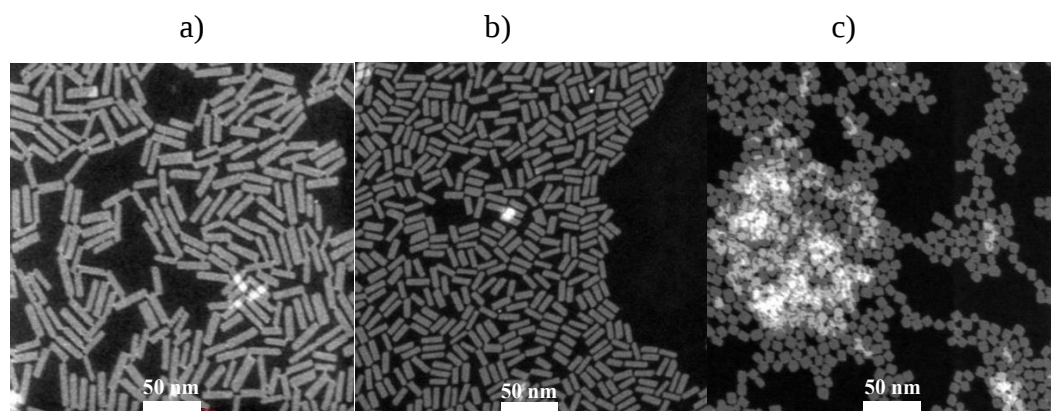


Figure 4.1: HAADF-STEM images of 4 ML NPLs synthesized using a) Cd acetate anhydrate, b) Cd acetate monohydrate, c) Cd acetate dihydrate.

On the other hand, depending on the variation of the amount of Cd acetate dihydrate, the shape can change from square to irregular forms (Figure 4.2, a-d). The major parameter changing here is the amount of water inside the crystal. However, although the NPLs in (d) are near to perfect square, they suffer from diminished chemical yield probably caused by excess Cd acetate. So, in terms of chemical yield, approximately 80-90 mg of Cd acetate is more appropriate under the given conditions. The same kind of experiments were carried out with the addition of pure water but was not successful in having the final square shape.

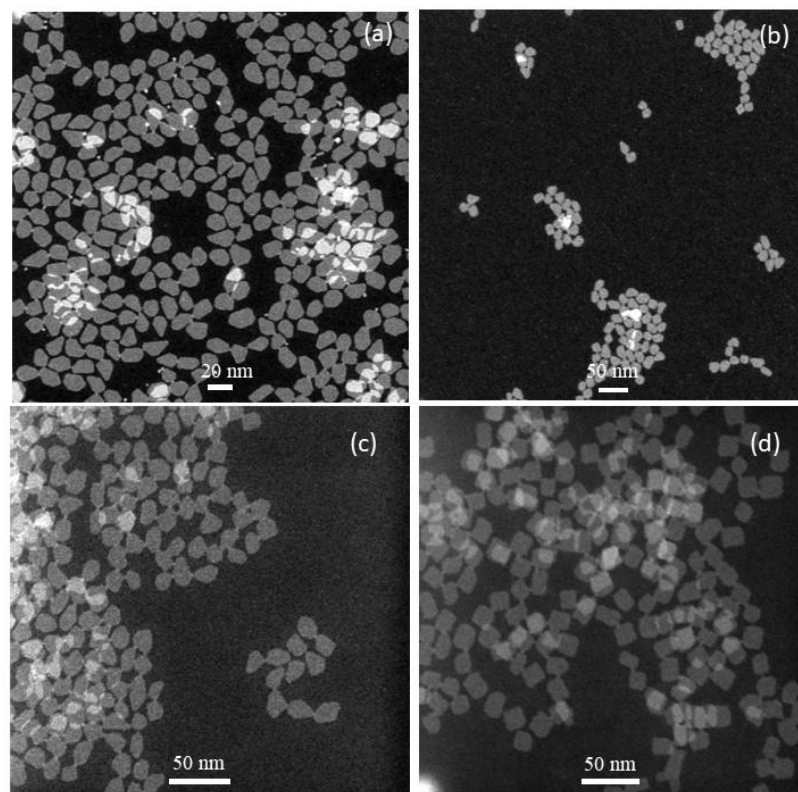


Figure 4.2: HAADF-STEM images showing influence of the Cd acetate amount on the final shapes of CdSe NPLs while all other conditions are kept the same: (a) 90 mg, (b) 100 mg, (c) 110 mg, and (d) 120 mg.

4.3. CdSe/ZnS core/colloidal atomic layer deposition shell synthesis

One of the techniques of synthesizing colloidal NPL heterostructures, as discussed before, is c-ALD. In the case of CdSe/ZnS core/c-ALD shell NPLs, a synthesis method differing from the typical literature pathway is used, mainly deviating in the chemicals used [19]:

- 1 mL of CdSe core NPL solution was cleaned with ethanol at least 3 times to ensure that surface ligands were torn off (assumption here is that if 100 μ L of CdSe NPLs are diluted by 3 mL of hexane, they have an optical density equal to 1 at 350 nm given the path-length of cuvette to be 1 cm).

- The final precipitate was dissolved in 3 mL hexane inside a 50 mL centrifuge tube and 3 mL NMF was subsequently added, which results in two-phase solution.
- In order to guarantee the complete coverage of the surface with S^{2-} ions, the excess of ammonium sulfide was used – around 100 μ L for each semi-layer, which results in transfer of the NPLs from non-polar hexane to polar NMF solution (color change was observed).
- The tube was placed upon a vortex for 5-10 minutes.
- Then the upper hexane layer was discarded.
- The precipitation was carried out by adding 0.5 mL acetonitrile and excess toluene and centrifuging for 3 min at 10,000 rpm.
- The process was repeated twice to ensure complete removal of ammonium sulfide and prevent any possible ZnS nanoparticle formation during the subsequent steps.
- The cleaned precipitate of CdSe/ S^{2-} was then dissolved in 2 mL NMF and combined with 2 mL 0.2 M Zn acetate dihydrate solution in NMF.
- After vortexing about 10-15 minutes, the cleaning procedure was conducted twice using toluene as a precipitating agent at 10,000 rpm for 3 min, this concluded the coating of monolayer ZnS.
- For any additional shell layers, the above-mentioned process was repeated.
- The NPLs were dispersed in hexane and 0.5 mL of OLA ligand was added.
- The sample was heated to promote capping of the surface with the ligand and was placed to the dark for at least 2 hours.
- Finally, the mixture was cleaned by ethanol at 6,000 rpm and hexane or toluene was added as a dispersant.

4.4. CdSe/ZnS core/hot-injection shell synthesis

One of the most important parts of our CdSe/ZnS project is the implementation of hot-injection shell growth around CdSe NPLs. As discussed in the previous chapters, in c-ALD growth method, although the number of layers can be manipulated easily, it is time-consuming and the synthesized hetero-NPLs, typically, are not chemically stable and do not have high quantum and chemical yield. That is the main motivation for which a new approach is necessary making the hot-injection shell growth a possible better alternative. The most important part is to choose suitable precursors, corresponding temperature range, optimum heating rates and the method of addition of the precursors. One of the most prominent sulphur sources is considered to be 1-octanethiol with the best operating temperature range being 280 – 330 °C. For this reason, the temperature was chosen to be 300°C. For Zn precursor Zn acetate was chosen to be as one of the best options. Below is the main procedure of the synthesis of CdSe/ZnS core/hot-injection shell NPL heterostructures. We developed:

- In a 50 mL round-bottom flask, 5 mL ODE, 36.5 mg Zn acetate (0.2 mmol), a certain amount of the 4 ML CdSe NPL and 500 μ L OA were introduced.
- As the CdSe NPLs were initially dissolved in hexane, in order to evaporate it, the solution was kept under vacuum (0.1–0.2 mbar) for about an hour at room temperature (RT).
- OA forms Zn oleate with the Zn acetate and produces water in the process. To remove this water and the remaining volatile chemicals, the flask was heated to around 90 °C and kept at that temperature for 30 min.
- Subsequently, the mixture was switched to Argon gas and 0.5 mL of OLA was added (It was shown before and observed that adding OLA at higher temperatures and long contact of CdSe NPLs with OLA can result in etching. Therefore, OLA was added when starting the reaction at a relatively lower temperature [29]).

- The final temperature was set to 300 °C.
- The octanethiol solution prepared beforehand (0.5 mL of 70 μ L octanethiol dissolved in 4 mL ODE) was injected dropwise using a syringe pump at the optimum temperature for the first drop being around 170 °C.
- The rate of the syringe pump was set to 10 mL/h at the beginning to ensure to start the growth and was dropped to 4 mL/h at 250 °C to sustain uniform growth on the surface and prevent fume formation at high temperature.
- The reaction for formation of the NPLs with the highest QY value is around 1 hour, which was confirmed experimentally (the results will be shown below). For that reason, the flask was preserved at 300 °C for an hour and then the reaction was quenched by cold water bath and 5 mL of hexane was added to increase dissolution.
- In order to clean the sample from possible side products the centrifugation at 6,000 rpm for 6 min was carried out. No precipitate was observed hinting that possibly no side products were formed.
- Finally, the sample was cleaned with ethanol once at 6,000 rpm for 6 min and was dispersed in hexane or toluene.

All of the above describes the synthesis of CdSe/ZnS thin – shell NPLs. They have 3.04 ± 0.3 nm thickness when the synthesis path described above is followed. However, it is well known that additional shell layers increase chemical stability and possibly QY of QDs [19], [28], [29]. To this end, another method producing thicker shells was used, which employed Zn oleate as one of the Zn precursors. Zn oleate was synthesized according to the literature with some modifications [30]:

- 915 mg Zn acetate (5 mmol), 3.15 mL OA and 16.85 mL ODE were loaded into a 50 mL round-bottom flask and the solution temperature was elevated to 110 °C under vacuum.
- The solution was maintained under this condition to get rid of the formed water,
- Later on, the temperature was increased to 150 °C to finalize synthesis of Zn oleate.

- Before cooling, the final mixture was kept at 100 °C for 10-15 min to evaporate the possible remnants of water.
- Finally, the solution was preserved at 80 °C under inert gas and used for the thick-shell NPL synthesis.

The synthesis of the thick-shell NPL, follows a similar route with some variations regarding the additional steps for the growth of the extra layers:

- Similarly, 5 mL ODE, 36.5 mg Zn acetate (0.2 mmol), a certain amount of the 4 mL CdSe NPL and 500 μ L OA were loaded to a 50 mL round-bottom flask.
- One-hour room-temperature degassing was undertaken to remove possible hexane left inside.
- To evaporate water possibly formed and hexane possibly left, the temperature was raised to 90 °C and kept there for about 50 min.
- The Schlenk line was then flushed with Argon gas, the temperature was set to 300 °C and 0.5 mL OLA was injected abruptly as soon as the temperature started to increase.
- As in the previous reaction, a syringe pump was utilized to add dropwise 4 mL ODE and 70 μ L octanethiol mixture prepared beforehand inside a glovebox (the rate was 4 mL/h for 30 min).
- Later, Zn oleate solution was introduced into the system in a dropwise manner (special care should be taken when adding with a hot syringe) and 5 min was allowed for the growth.
- The addition of the S precursor was resumed at a 5 mL/h rate.
- After allowing a 40 min period, the reaction was quenched with cold water and 5 mL hexane was injected.
- The cleaning procedure was identical to the thin-shell NPLs yielding almost no side-products.

4.5. CdS core synthesis

When it comes to CdS core NPL synthesis, it is harder to control the final shape, as the as-synthesized NPLs tend to fold due to huge lateral dimensions (in contrast to CdSe). Also, even when 4 ML CdS NPLs are synthesized and purified via size-selective precipitation, it is very challenging to achieve pure 4 ML without some tiny amounts of 3 ML and 5 ML CdS NPLs being present. In fact, it is not possible to clean as much as CdSe because ligands are torn away from the surface much easier. The method used in this work for the synthesis is similar to the literature except for some minor modifications [14]:

- 0.1 M of S-ODE precursor was prepared beforehand by heating the corresponding amount of S inside ODE at 120 °C (e.g., for making 20 mL 0.1 M solution, 64 mg of S was dissolved in 20 mL ODE under ambient conditions).
- 217 mg of Cd acetate dihydrate, 2 mL of 0.1 M S-ODE solution, 300 mL of OA and 10 mL of ODE were placed inside a 50 mL round-bottom flask.
- Room-temperature Argon bubbling was used as the most efficient way for removing contaminants for at least 30 mins (in fact, degassing at room temperature results in a low chemical yield of 4 ML NPLs and the percentages of 3 and 5 ML CdS NPLs are higher),
- Then, the temperature was increased to 240-250 °C at a rate of about 23 °C/minute and as the solution reaches the desired value, it is kept there for at least a minute: the longer the reaction period the bigger the produced lateral sizes.
- Finally, the reaction was quenched by placing carefully in room-temperature water under the flask and injecting 5 mL hexane after the solution reached 120 °C.

It should be noted that, unlike the previously reported value of 260 °C for the synthesis temperature, when the temperature was kept below 240 °C or above 250 °C 3 and 5 ML CdS NPLs were formed. The deviation may be explained by setting values unique to each heating mental system.

As far as the cleaning process was concerned, ethanol usage was avoided here. Ethanol is more polar compared to acetone. That is why initially weak bonds between CdS semiconductor NPL and carboxylic acid ligands are broken easier with ethanol. As a result, coagulates were formed. Acetone, on the other hand, is weaker, so, chemical and quantum yields are less affected. At first, centrifugation was carried out at 4,000 rpm for 6 min to precipitate contaminants. Some bigger and less stable NPLs were precipitated, as well.

4.6. CdS/CdSe core/crown synthesis

The recipe used for CdS/CdSe core/crown heteronanoplatelets synthesis differs from the literature [31]. There are two different Se precursors used: ODE-Se and TOP-Se. For the preparation of 0.1 M ODE-Se solution, 15 mL ODE and 118.5 mg Se were loaded to a 50 mL flask. It was degassed for 1-2 min and then switched to Argon gas. It was heated to 90 °C and kept at it for 1 hour. Then the temperature was raised to 120 °C and maintained there for half an hour. Subsequently, it was kept at 150 °C and then 180 °C for half an hour each. It was cooled and kept for later usage. To make 0.1 M TOP-Se solution, 0.1 mmol of Se was added to 1 mL TOP inside a glovebox because TOP is flammable in air. It was mixed at 50 °C to dissolve faster. Cd precursor was also made separately beforehand as described in CdSe/CdTe core/crown synthesis in Chapter 2.2. After producing necessary precursors and synthesizing CdS NPLs the following steps were implemented:

- CdS NPLs (amount equalling to 0.5-1 O.D.), a certain amount of crown precursors (typically 500 µL), a little amount of OA (usually between 50-75 µL) and 5 mL ODE were added to a 50 mL round-bottom flask.

- The mixture was degassed at room-temperature for at least 1 hour to remove any hexane and contaminant volatile substances, which might be present. Then it was degassed at 90 °C for 20-25 min.
- Meanwhile, TOP-Se solution was prepared for injection with a syringe pump: inside the glovebox, 60 μ L of 0.1 M TOP-Se solution was mixed with 2 mL ODE and 0.5 mL of it was taken by a 20 mL syringe avoiding any bubbles inside.
- The system of the reaction was changed to Argon gas and temperature was set to 240 °C.
- The first drop of Se precursor was set to be around 180 °C and the rate was 8 mL/h.
- The reaction vessel was kept at 240 °C for about 8-10 min.
- Then the reaction was quenched by cold water bath and 5 mL of hexane was poured for better dissolution.
- The cleaning was performed by using both acetone and ethanol after the first precipitate was disposed of by 2,000 rpm centrifugation.

Depending on the quantity of Cd crown precursor employed and time period of the reaction, the formed core/crown structure may become more red-shifted. Better results were obtained when TOP-Se precursor was used instead of ODE-Se. This can be caused by better dissolution of Se inside TOP and its assistance when attaching to CdS NPL surface.

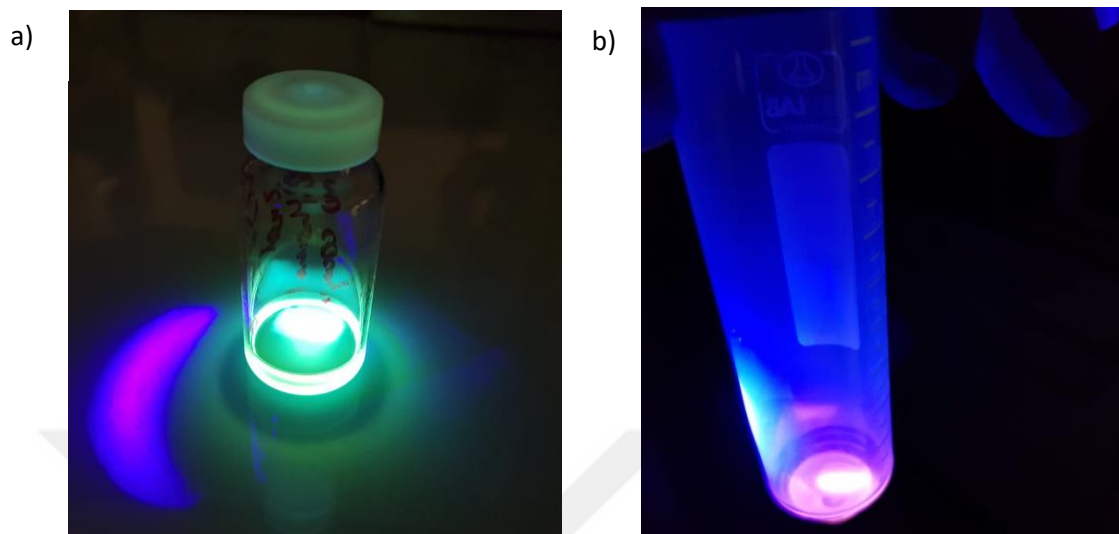


Figure 4.3: Photographs of a) our CdS/CdSe inverted NPLs and b) our CdS/CdSe/S²⁻ core/crown/ c-ALD semi-layer NPLs when illuminated with a UV light source in the dark.

4.7. CdS/CdSe/ZnS core/crown/colloidal atomic layer deposition shell synthesis

Although attempts to synthesize CdS/ZnS core/hot-injection shell were not successful, CdS/CdSe/CdZnS core/crown/c-ALD shell NPL synthesis pathway led to good results in terms of the final shape of the structures. Then NPLs were synthesized in a way similar to CdSe/ZnS core/c-ALD shell heterostructures with some deviations [19]:

- The solution of CdS/CdSe NPLs (in an amount equal to 0.5 O.D.), synthesized by the method described in Chapter 4.6 was cleaned with ethanol for 1-2 times

to ensure that surface ligands were torn off (our assumption here was that if 100 μL of CdS/CdSe NPLs were diluted by 3 mL of hexane, they have an optical density equalling to 1 at 350 nm wavelength given the path-length of cuvette to be 1 cm).

- The final precipitate was dissolved in 1mL hexane inside a 50 mL centrifuge tube into which 20 μL NH_4S (in excess quantity) and 400 μL NMF was added. This resulted in two-phase solution and color change began to be observed.
- The tube was placed in a vortex mixer for 5-10 minutes.
- Then the upper hexane layer was discarded.
- The precipitation was carried out by using 1/1/2 volume/volume (v/v) sample/acetonitrile/toluene solution and centrifuging at 4000 rpm for 4 min.
- Subsequently, the second precipitation was implemented by using 1/1/4 v/v sample/acetonitrile/toluene solution and centrifuging at 4000 rpm for 4 min.
- This ensured complete removal of ammonium sulfide, preventing any possible ZnS nanoparticle formation during the subsequent steps.
- The cleaned precipitate of CdS/CdSe/S^{2-} was then dissolved in 400 μL NMF and combined with 400 μL of 0.3 M Zn acetate dihydrate and Cd acetate dihydrate solution in NMF (the solution was prepared beforehand by dissolving 1.5 mmol Cd acetate dihydrate and 1.5 mmol Zn acetate dihydrate in 10 mL NMF).
- After vortexing about 4-5 minutes, the cleaning procedure was conducted twice using toluene in a 1/4 v/v ratio at 5,000 rpm for 5 min.
- This completes the coating of monolayer $\text{Cd}_{0.5}\text{Zn}_{0.5}\text{S}$.
- For any additional monolayers of the shell growth the process described above was repeated.
- The NPLs were dispersed in hexane and 0.4 mL of OLA and 0.4 mL OA ligands were added.
- The sample was heated to promote capping of the surface with the ligand and was placed in dark for at least 2 hours.

- Finally, the mixture was cleaned by ethanol at 6,000 rpm and hexane or toluene was added as a dispersant.



5. Characterization of nanoplatelet heterostructures coated with crown and shell layers

This chapter is partly based on our publication “Highly Stable, Near-Unity Efficiency Atomically Flat Semiconductor Nanocrystals of CdSe/ZnS Hetero-Nanoplatelets Enabled by ZnS-Shell Hot-Injection Growth,” [2]

5.1. Optical characterizations

5.1.1. CdSe/ZnS hot-injection NPLs

When thinking about possible application areas of fluorescent NCs, the first properties to be analysed and to be tuned are the optical ones. In the case of CdSe/ZnS hot-injection NPLs synthesized in this thesis work, they show outstanding tunability and narrow FWHM compared to the c-ALD counterparts:

a)

Absorbance

b)

Absorbance

Figure 5.1: Absorption and PL emission spectra of CdSe/ZnS hot-injection a) thin-shell and b) thick-shell samples.

The typical PL emission peak for thin-shell NPLs appears at around 600-602 nm, with a slight shift depending on fluctuations of the reaction conditions and the Schlenk line system used. From Figure 5.1, it is obvious that for thin-shell NPLs the PL peak is at 600 nm and their FWHM is 23 nm, which is quite remarkable for the core/shell NPLs. On the other hand, the thick-shell NPL peak is not as much red-shifted, which is located at around 610 nm, with a FWHM of 27 nm. QY, FWHM and PL emission peak data of various thin-shell and thick-shell NPL samples were collected which are presented in Table 5.1.

Table 5.1: QY (%), FWHM and emission peak (nm) values of 10 distinct thin-shell and thick-shell NPL samples.

Thin-shell NPLs			Thick-shell NPLs		
QY (%)	Emission peak (nm)	FWHM (nm)	QY (%)	Emission peak (nm)	FWHM (nm)
88	606	22	92	606	25
87	606	23	95	604	25
82	625	26	100	600	23
94	602	25	98	619	29
92	606	26	91	613	28
83	602	24	93	608	25
81	625	25	98	606	24
81	633	27	87	592	25
80	581	24	95	600	23
77	588	24	98	613	26

The highest and lowest values of QY (%), PL emission peak (nm) and FWHM (nm) do not deviate considerably from each other in both cases of the thick- and thin-shell NPLs, which points out the reproducibility of the recipes.

To understand the growth process of CdSe/ZnS thin-shell NPLs, aliquots were taken and were analysed using the optical instrumentation. The whole reaction span was 3 hours with certain time intervals (Figure 5.2).

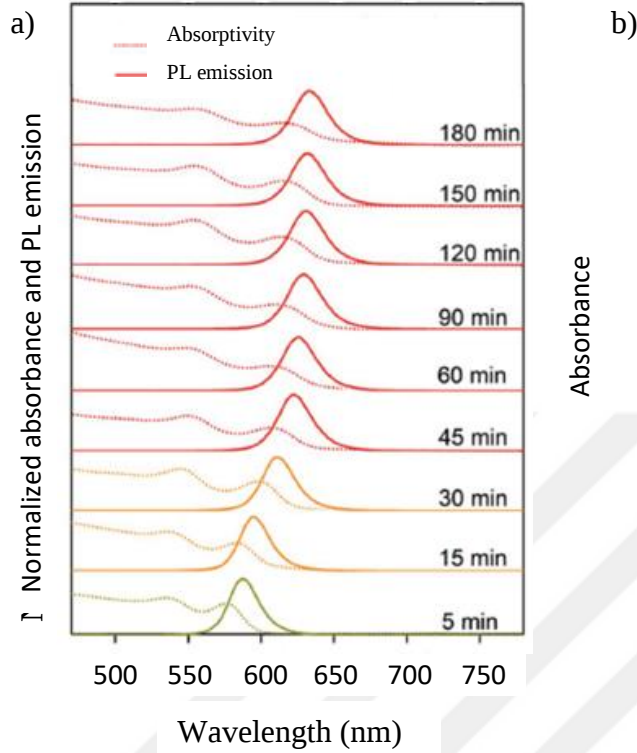


Figure 5.2: a) Absorption and PL emission of the aliquots of the thin-shell sample taken in certain time intervals and b) those of the final cleaned sample. a) Reprinted with permission from [2].

From Figure 5.2, it can be concluded that right from the first few minutes of the reaction ZnS shell was being formed as the red-shifted peaks were then observed. The amount of red-shift from the 510 nm PL peak of 4 ML core at the 5th, 30th, 60th, 120th and 180th min of the reaction was 75 nm, 98 nm, 112 nm, 116 nm and 120 nm, respectively. The CdSe/ZnS core/shell NPLs exhibit type-I electronic structure, which results in electron and hole pair being confined to the same semiconductor layer. The red-shift can be attributed to the change of the effective dielectric constant again coming from heterostructure. Another crucial change observed during the 3-hour period of reaction was a slight increase of FWHM from 22 nm in the first 30 min to 26 nm at the 90th min and to 29 nm at the end of 3 hours. The reason behind it is the enhanced exciton-phonon coupling inside the formed heterostructures, the large lattice

mismatch between core CdSe and ZnS shell materials, which causes lattice strain and finally, rise in the number of NPLs having different thicknesses [20][32]. At first, during the first 30 min, the QY rose from 4% at the 2.5th min to 60% at the end. This suggest that the walls of the formed heterostructures were passivated with ZnS as the shell growth took place. At the end of 60 min, the QY reached its maximum value of 77% and then started gradually decreasing. At the end of three hours, it fell to 50%, which implies that extended growth time results in possibly the formation of more defect sites and increased strain as was discussed before.

Another optimization procedure for the synthesis recipe was conducted by varying the amount of OA and OLA, one at a time (Figures 5.2 and 5.3).

Table 5.2: QY (%), PL emission wavelength (nm) and FWHM (nm) values based on the OA amount used (500 μ L OLA is used in each case).

Volume of OA (μ L)	QY (%)	Wavelength (nm)	FWHM (nm)
0	75	592	28
500	85	602	24
1000	56	625	21

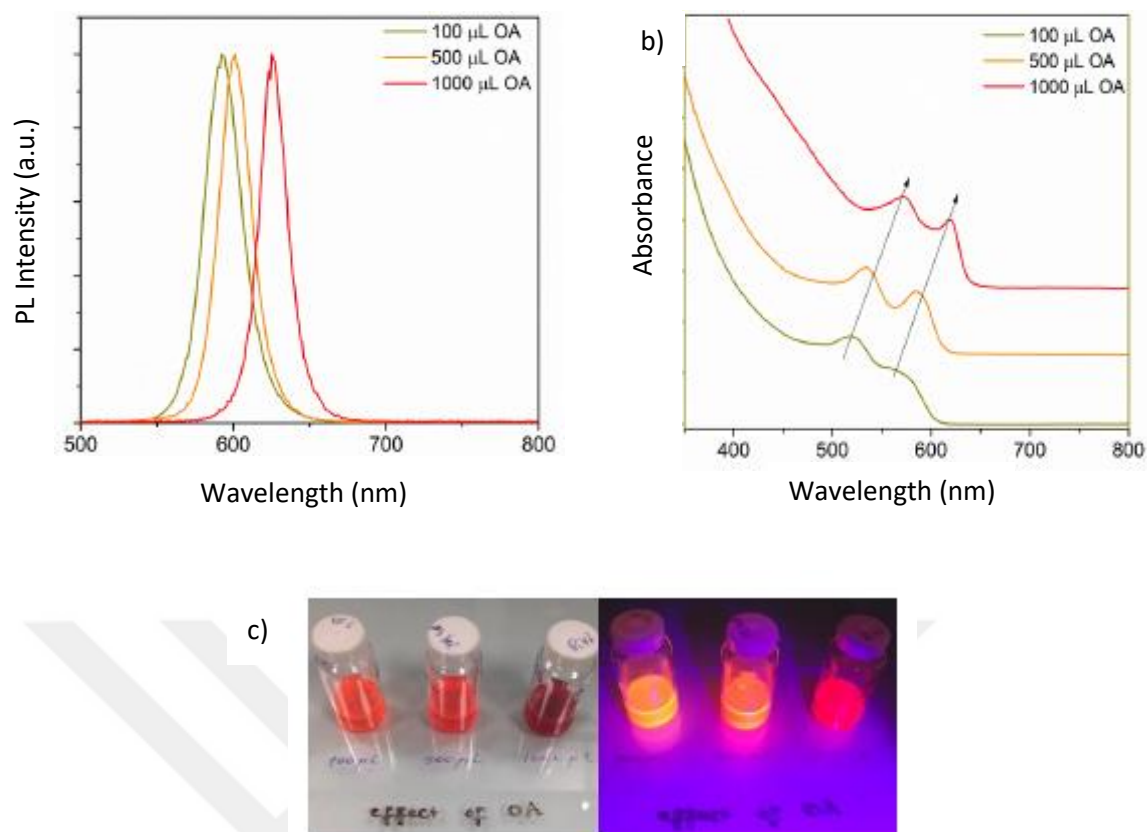


Figure 5.3: a) PL emission and b) absorbance spectra and images of CdSe/ZnS hot-injection core/shell samples based on the variation of OA amount in the synthesis. Reprinted with permission from [2].

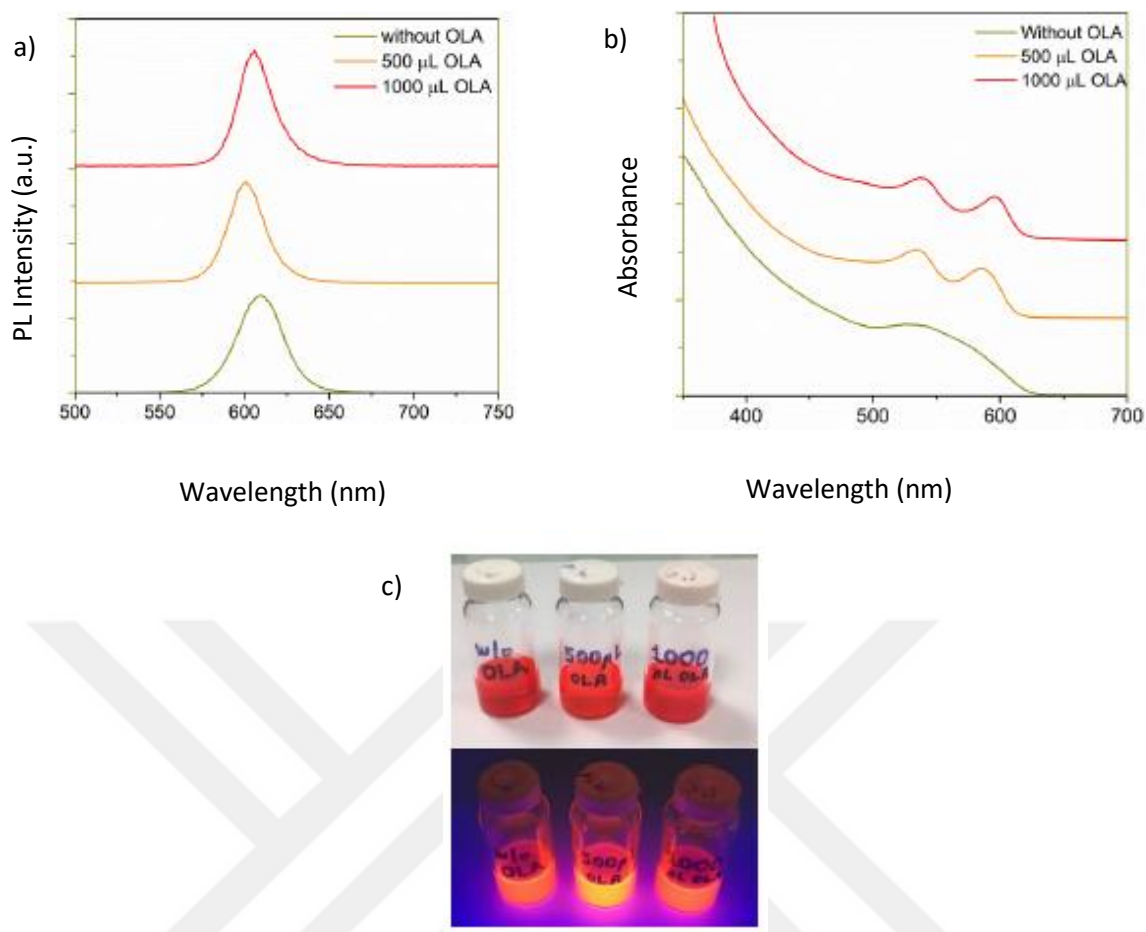


Figure 5.4: a) PL emission and b) absorbance spectra and c) images of CdSe/ZnS hot-injection core/shell samples based on the variation of OLA amount in the synthesis. Reprinted with permission from [2].

Table 5.3: QY (%), PL emission wavelength (nm) and FWHM (nm) values based on the OLA amount used (500 μ L OA is used in each case).

Volume of OLA (μ L)	QY (%)	Wavelength (nm)	FWHM (nm)
0	69	610	31
500	85	602	24
1000	82	605	24

To compare the synthesis of CdSe/ZnS hot-injection NPLs with the previously reported ones, CdSe/ZnS core/shell NPLs were synthesized also by using c-ALD method and their optical properties were compared to our hot-injection shell

heterostructures [20], [33]. The synthesis was done as discussed before at room temperature and the number of shell layers was tuned from 1 to 5. The optical spectra of our synthesized c-ALD heterostructures are presented below (Figure 5.5).

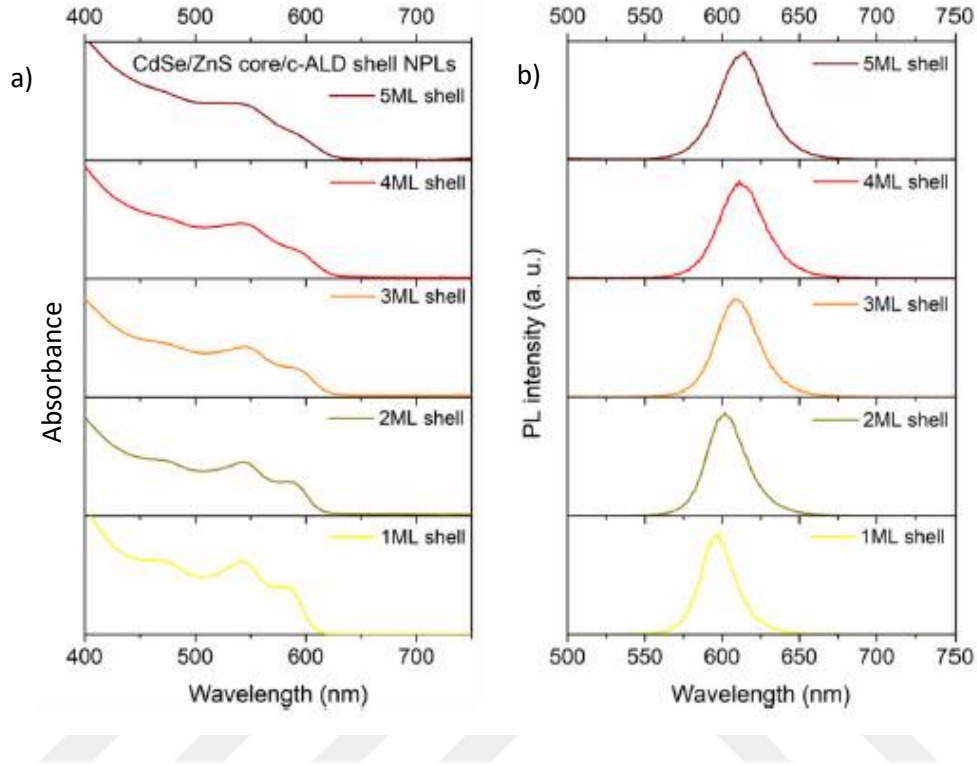


Figure 5.5: (a) Absorbance and (b) PL spectra of CdSe/ZnS core/c-ALD shell heterostructures synthesized at our laboratory.

From Figure 5.5, it is concluded that FWHM of the samples ranged from 25 to 35 nm and their peak emission wavelength shifted from 595 to 614 nm when the number of layers increased from 1 to 5 ML. Unlike their hot-injection counterparts, the c-ALD shell grown NPLs did not dissolve after 5 times of cleaning with ethanol solvent due to instability and low concentration of ligands on the NPLs' surface. The QY values of these c-ALD shelled NPLs were considerably weaker compared to our hot-injection shell grown NPLs, which stems from better surface passivation in the core of the HI growth.

5.1.2. CdS NPL heterostructures

The CdS NPLs previously synthesized in the literature exhibit absorption maximum at around 410 nm, while their PL emission peak is at approximately 420 nm (FWHM 12 nm), accompanied by very broad trap emission peak around 646 nm [5]. The light hole and heavy hole transitions of the absorption spectra are located close to each other. The QY is very low (around 4%).

For the CdS NPL samples synthesized in this thesis work, the trap emission was reduced comparatively to a negligible amount as can be seen from Figure 5.6. When it comes to purity, unfortunately, like in the most of literature data, 3 and 5 ML NPLs were not completely removed [14]. However, because the reaction yield was high, this can be reduced to a negligible amount.

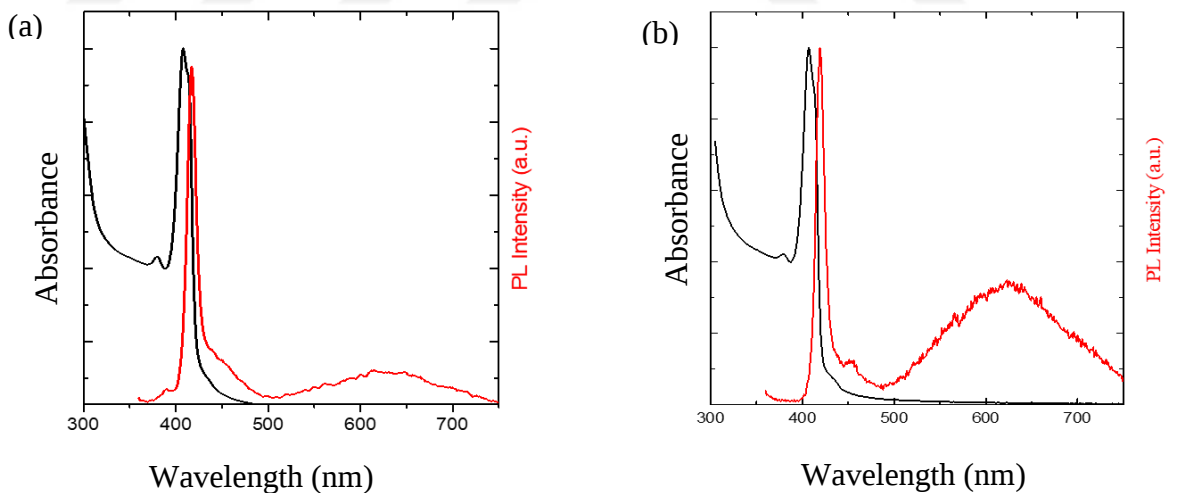


Figure 5.6: (a) Absorption and (b) PL emission spectra examples of two CdS NPL samples synthesized at our laboratory.

In the case of CdS/CdSe core/crown NPLs, the results are promising because of the increased QY, almost completely removed trap emission and enhanced chemical stability, which allows the cleaning to be implemented by a stronger polar solvent - ethanol. The amount of added Cd and S precursors determine the thickness of the

crown layer formed and, hence, the red-shift of the final structure. As it can be inferred from Figure 5.7, the absorption spectrum did not deviate considerably from that of CdS NPLs. This indicates that no extra layers covering the sample NPLs in the thickness direction or any shell is formed. On the other hand, the red-shift of PL emission suggests a substantial change in the dielectric medium due to the lateral extension. The fact that adding more precursors results in narrower crown formation means that excess reactants are depleted for other purposes.

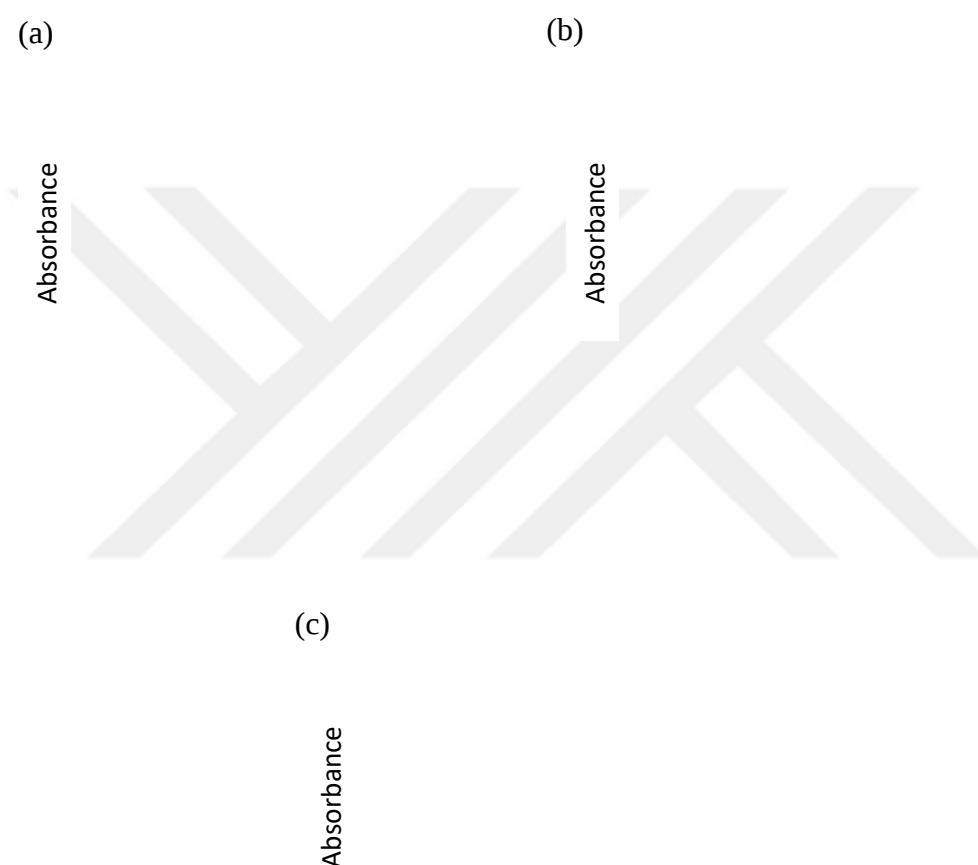


Figure 5.7: Absorption and PL emission spectra of CdS/CdSe NPLs for the cases (a) when 1 mL Cd crown and 1 mL of TOP-S precursor, (b) when 500 μ L Cd crown and 0.5 mL TOP-S precursor, and (c) were used.

Although the experiments performed for the growth of hot-injection shell layer were not successful, the results indicate that it is possible to find a chemical route for hot-injection shell coating. After coating ZnS c-ALD and unfolding the CdS NPLs, the hot-injection was undertaken to obtain NPLs with the optical properties shown in Figure 5.8. Though there is trap emission observed implying that the chosen pathway, temperature and/or precursors did not provide the best combination for the shell growth, there is a shell layer formed with PL emission at around 500 nm. This means that there is still a potential for using the hot-injection shell coating for CdS heterostructures.

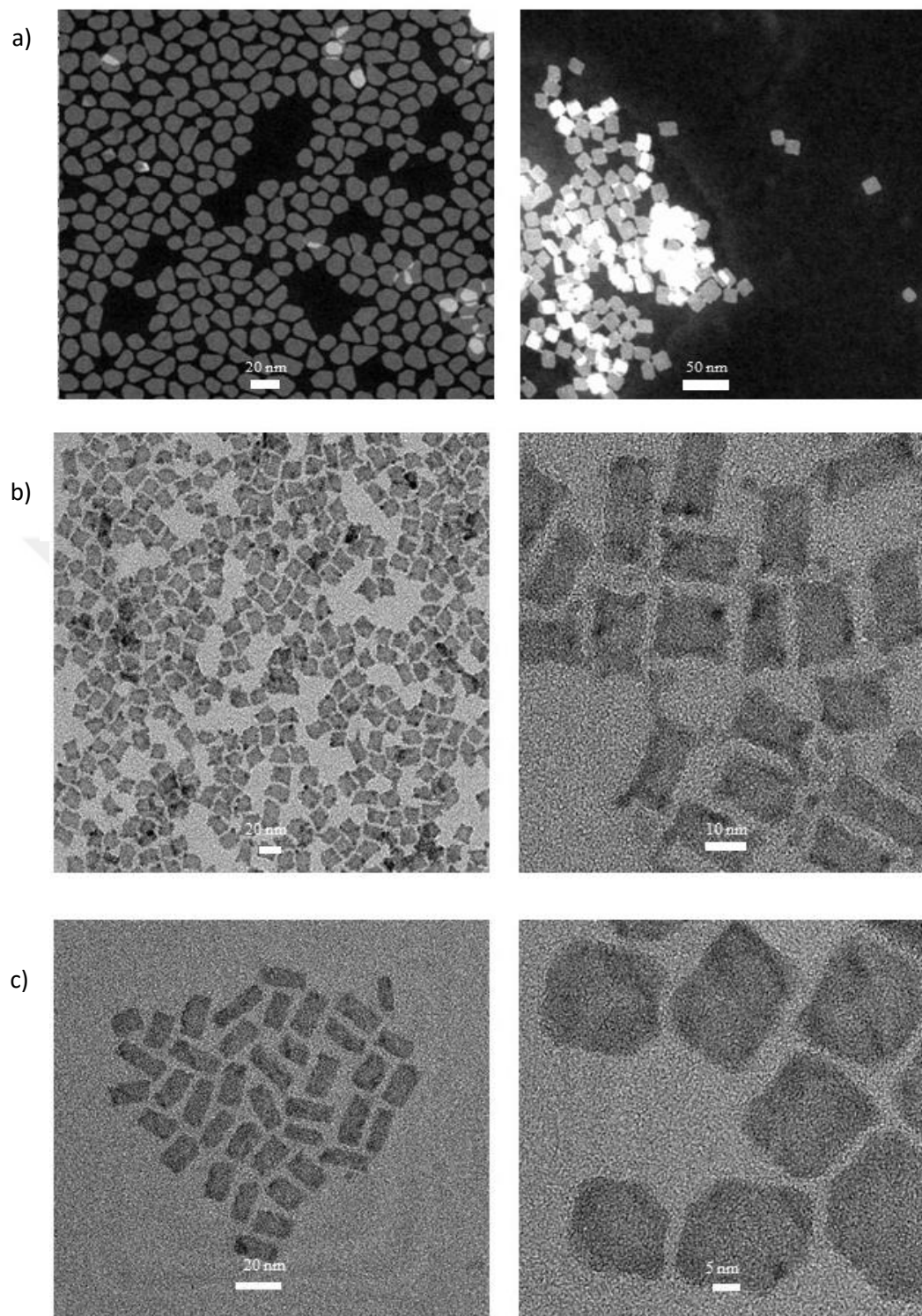


Figure 5.8: Absorption and PL emission spectra of CdS/c-ALD ZnS/hot-injection ZnS shell NPLs.

5.2. Structural characterizations

5.2.1. CdSe NPL heterostructures

One of the key means to understand and improve the chemical synthesis of any nanocrystal is to employ structural characterization by using various microscopes, XRD and XPS instrumentation. In this thesis work, special effort was spent on acquiring high-quality TEM images, XRD and XPS data for the following elaborate analyses. The CdSe core NPLs used for the synthesis of our core/shell NPLs mostly had square shapes and it was observed that using irregular shaped NPLs does not produce desired results. However, when square shaped CdSe NPLs were used, the resultant CdSe/ZnS core/hot-injection shell NPLs acquired square or rectangular shapes and possessed enhanced optical properties. High-resolution TEM (HR-TEM) images in Figure 5.9 prove this point. On the other hand, thick-shell NPL heterostructures also tended to yield rectangular shapes with elongated tips which were found to consist of ZnS based on EDS analysis. In some images taken from different synthesis batches, it was observed that some NPLs lacked those tips and some parts of them were detached (Figure 5.10). To measure the thickness of the synthesized NPLs, TEM images containing free-standing NPLs were captured (Figure 5.11). The thickness of thin-shell samples was measured to be around 3.04 ± 0.3 nm, while it was increased to 4.03 ± 0.43 nm in thick-shell samples.



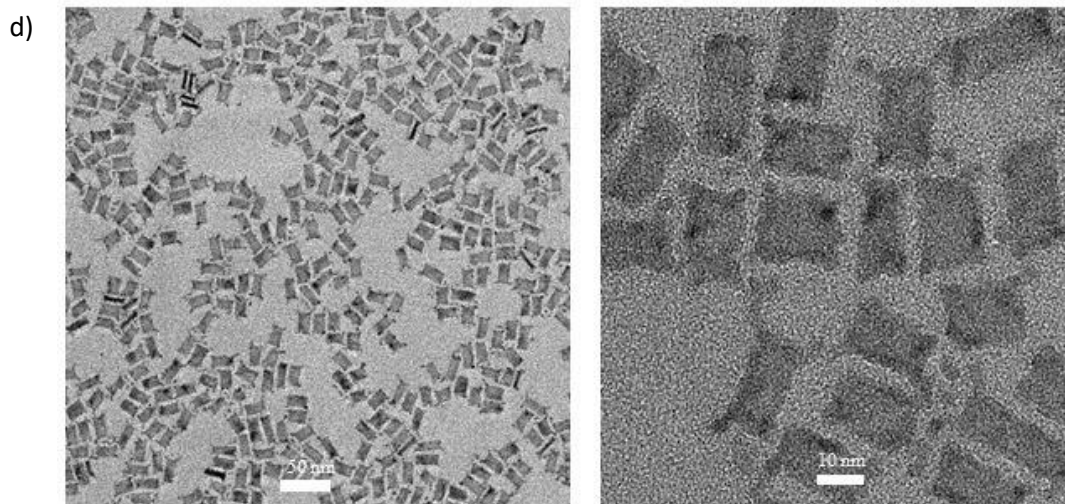


Figure 5.9: a) HAADF-STEM images of CdSe NPLs with irregular (left) and square (right) shapes; HR-TEM images of CdSe/ZnS core/hot-injection thin-shell NPLs synthesized using b) irregular shaped CdSe core NPLs and c) square shaped CdSe NPLs; and d) HR-TEM images of thick-shell CdSe/ZnS NPLs.

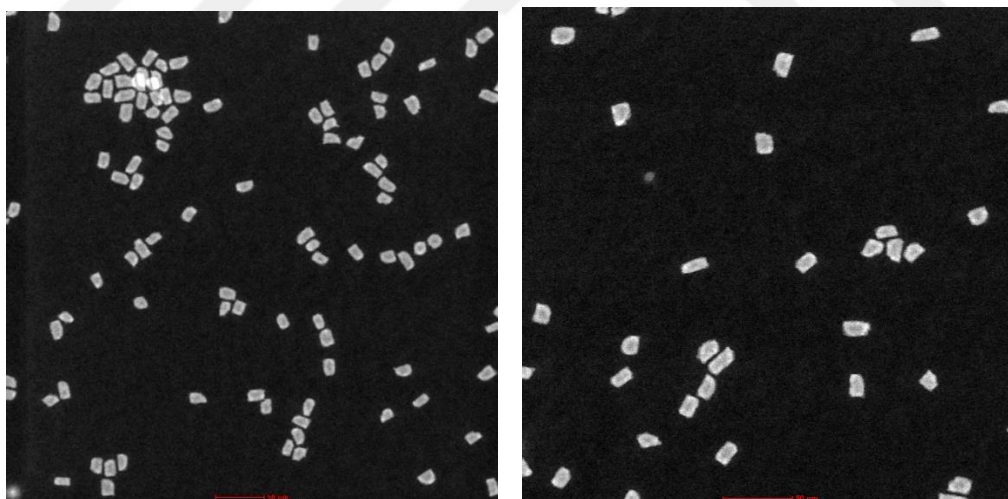


Figure 5.10: HAADF-STEM images of CdSe/ZnS core/hot-injection thin-shell NPLs where it is evident that their tip parts are detached during synthesis, cooling or cleaning.

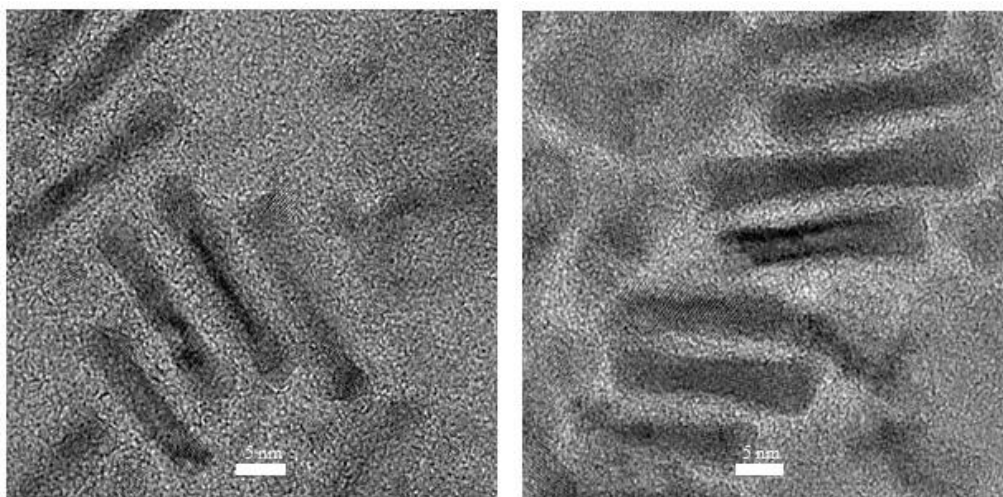


Figure 5.11: HR-TEM images of free-standing CdSe/ZnS core/hot-injection thick-shell (left) and thin-shell (right) NPLs.

Apart from TEM, the SEM of focused ion beam (FIB) instrument was employed for the structural analysis. Due to its worse resolution, although a minimum of 100 nm was resolved, 3D shapes were visible by using this method (Figure 5.12)

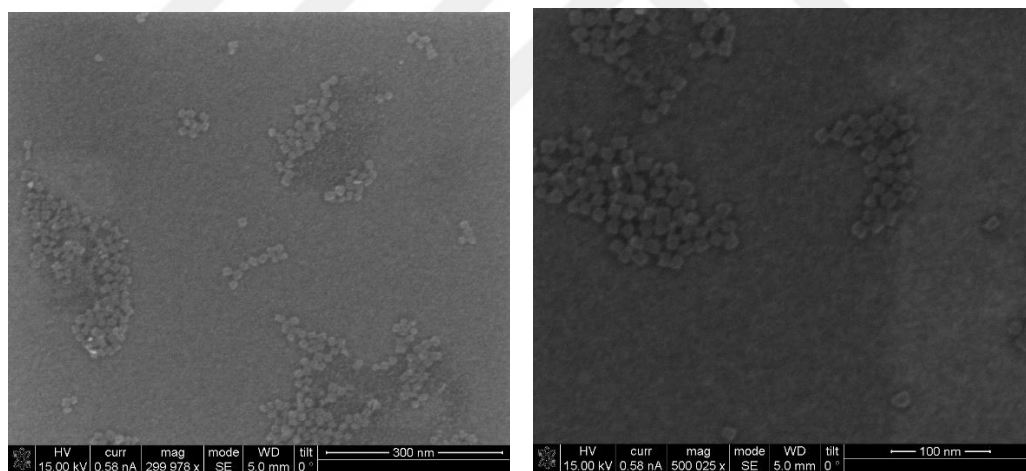


Figure 5.12: SEM images of CdSe NPLs.

One of the other important studies regarding the structure of the NPLs was based on XRD analysis. This was carried out by X-ray diffractometer for powder (XRD- MPD) instrument (Panalytical Xpert Pro MPD). The NPL samples for XRD analysis were prepared by drop-casting NPLs onto glass slides along a line and placing this glass inside the XRD chamber in the direction by which this line will coincide with the laser. The collection range after several trials were determined to be between 20° to 70° (Other parts of the spectra lacked any meaningful data). Our XRD measurements are

shown in Figure 5.13. For comparison purposes, XRD peaks of zinc blende bulk CdSe (black vertical lines at the bottom) for the core and ZnS (red vertical lines at the top) for the shell are also depicted on the graph. It is comprehended from this graph that the more shell layers are formed, the greater shifts from CdSe bulk XRD peaks towards ZnS bulk XRD peaks are. Alongside with other data, this XRD spectrum shows that, indeed, the shell growth was successful and new crystalline structure formed. Also, it can be seen that, in general, the largest peak belongs to CdSe/ZnS hot-injection thick-shell NPLs. It might mean that they have the highest value of crystallinity when compared to their c-ALD counterparts. Their peaks are sharper at (111) plane, also indicating more crystalline formation.

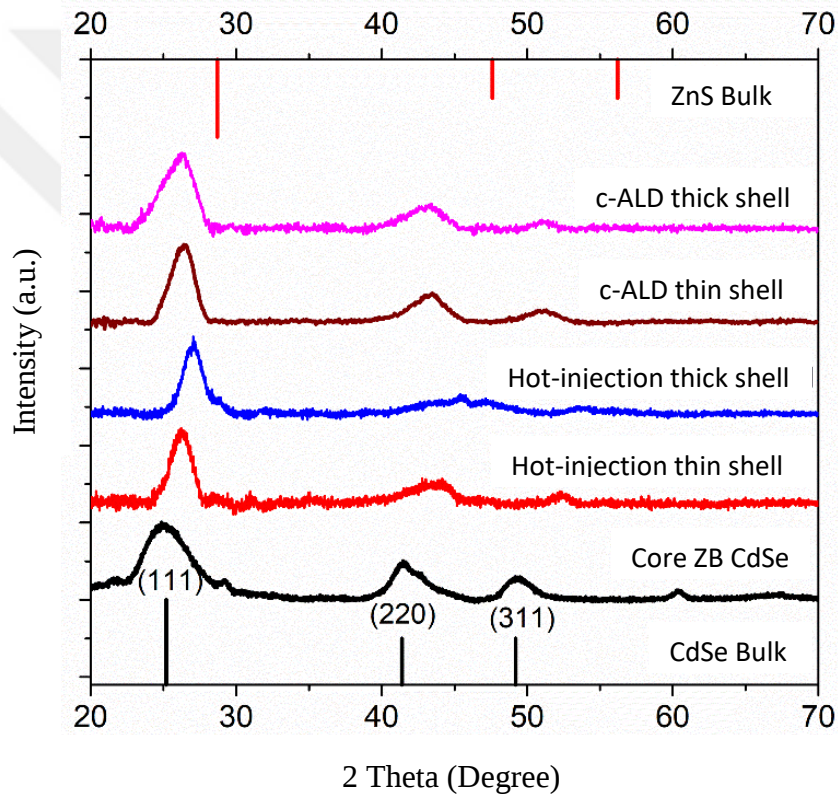


Figure 5.13: XRD data of our CdSe core, CdSe/ZnS core/c-ALD shell and CdSe/ZnS core/hot-injection shell NPLs. Reprinted with permission from [2].

Another crucial data set for the structural analysis was obtained by using XPS instrument. A full set of the collected data is given in Appendix A1. This data was used to confirm the presence of certain elements in the NPL structures alongside EDS analysis.

5.2.2. CdS NPL heterostructures

One of the worst features about CdS NPLs is their size and shape. In most cases NPLs with a rather large surface area, especially in contrast to their thickness, were produced. This causes surface stress and NPLs roll like scrolls (Figure 5.13). Due to this tendency, although CdS NPLs have promising potential applications, particularly in lasing, CdS NPLs are extremely hard to work with. It is necessary to unfold the structure and later cover it with some shell coating for chemical protection and reduction of defect sites. In this work, the solution came in the form of crown and c-ALD shell coating combined together. As it can be inferred from Figure 5.14 the final structures, though differing in sizes and shapes, were all unfolded and detached from each other.

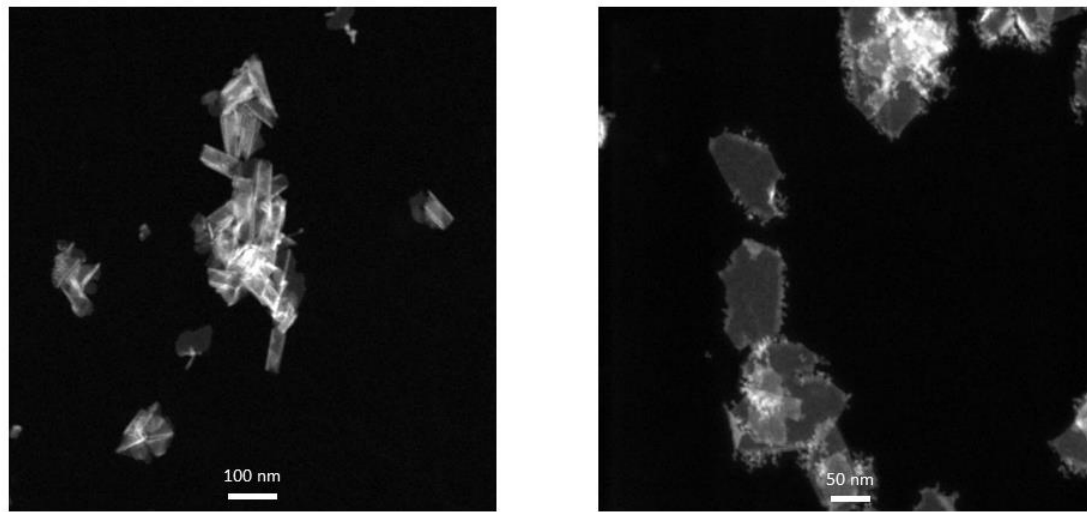


Figure 5.14: TEM images of CdS NPLs synthesized at our laboratory: the left image shows the typical rolled CdS NPL structures and the right one depicts their actual shape when unrolled.

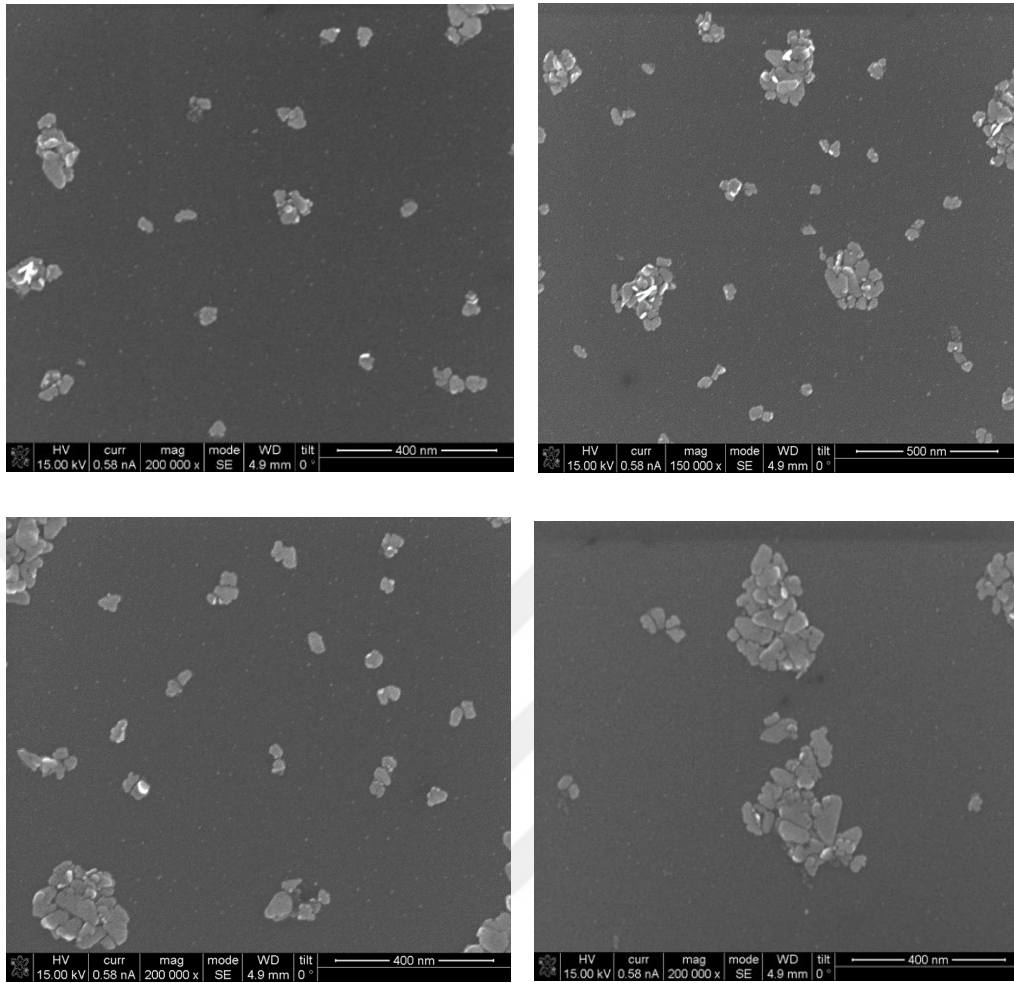


Figure 5.15: SEM images of CdS/CdSe/CdZnS core/crown/c-ALD shell NPLs synthesized at our laboratory.

5.3. Aging tests

5.3.1. Thermal aging

One of the key characteristics defining potential application areas of semiconductor NCs including NPLs is their level of thermal endurance. For instance, when LEDs are designed, the preferred temperature elevation endurance for the material to be used is at least 400 K. To test the thermal durability of semiconductor NCs some thermal tests should be carried out encompassing, preferably, tests both under ambient conditions and under inert gas. As discussed in Chapter 3.3, Rowland et al. used static and

transient spectroscopies as well as in-situ TEM to analyze thermal properties of CdSe and CdSe/CdS NPLs [28]. The temperature they used in their tests ranged from 300 to 700 K under vacuum. Drop-cast films of CdSe NPLs, CdSe/3 ML CdS core/shell NPLs and CdSe/6 ML CdS core/shell NPLs synthesized via c-ALD method were employed. Their findings indicated whether the changes observed in PL intensity are of the reversible (physical modifications) or of irreversible (chemical modifications) nature.

The main reason behind PL quenching at elevated temperatures is the hole trapping. In Rowland et al., for static spectroscopy analysis, the sample drop-cast films were placed inside an oven that was degassed beforehand. Subsequently, time-integrated PL spectra were collected using a 405 nm diode laser. The PL intensities were found by integrating collected spectra or by following the PL maximum. By employing temperature-dependent TEM images, the effect of the temperature elevation was directly observed. It was concluded that the shape and morphology was preserved until 560 and 630 K, respectively. However, above those values, the obvious deterioration happened, which lead to change in confinement of NPLs and substantial loss of optical properties.

In this thesis work, the thermal properties were tested straight after our synthesis under inert gas. The setup was prepared as shown in Figure 5.17. The thin-shell and thick-shell grown CdSe/ZnS core/shell NPL samples were kept inside the flask right after the reaction was completed and they were in situ heated up to 525 K. The changes in the PL intensity were detected by OceanOptics MAYA 2000 model spectrophotometer in 25 K steps. In order to obtain better results, special attention was paid to record values with 0.1 K fluctuations. The lamp employed was stabilized to a certain spot to avoid any deviations in the exposure of the light. At first, cooling was allowed right after the reaction, the sample was purified through precipitation by centrifugation. Then, it was dissolved into 10 mL ODE and poured into a 50 mL flask for testing. The spectrophotometer was connected to the computer and the PL cable was set to the appropriate angle. The temperature was set to 500 K to avoid overshooting. The PL intensity values were recorded starting from 300 K. From Figure 5.16 it can be inferred that the final FWHM shift of PL spectra for the thin-shell CdSe/ZnS core/shell samples was 30 nm and for the thick-shell sample it was 29 nm. One of the key points to note is that, at 400 K, the thin-shell samples preserved 42% of their initial PL intensity, while the thick-shell samples sustained 65% of their initial PL intensity values. This is

crucial in the sense that typical LED and laser applications reach up to 400 K while operating. After the temperature reached 525 K, 15% of the initial PL intensity of the thin-shell sample was preserved and 37% of the initial PL intensity of the thick-shell one was conserved. Therefore, these NPLs have potential applications not just in LEDs, but in other fields as well. One of the key observations is that, after the solutions cooled by themselves, when the heating source was turned off, the beginning PL was restored to almost 100% for the thick-shell sample, making it quite unique in its class of NPLs and among semiconductor NCs in general. The thin-shell sample's PL recovered up to 74%, which is also quite remarkable. These cycles are depicted in Figure 5.16.

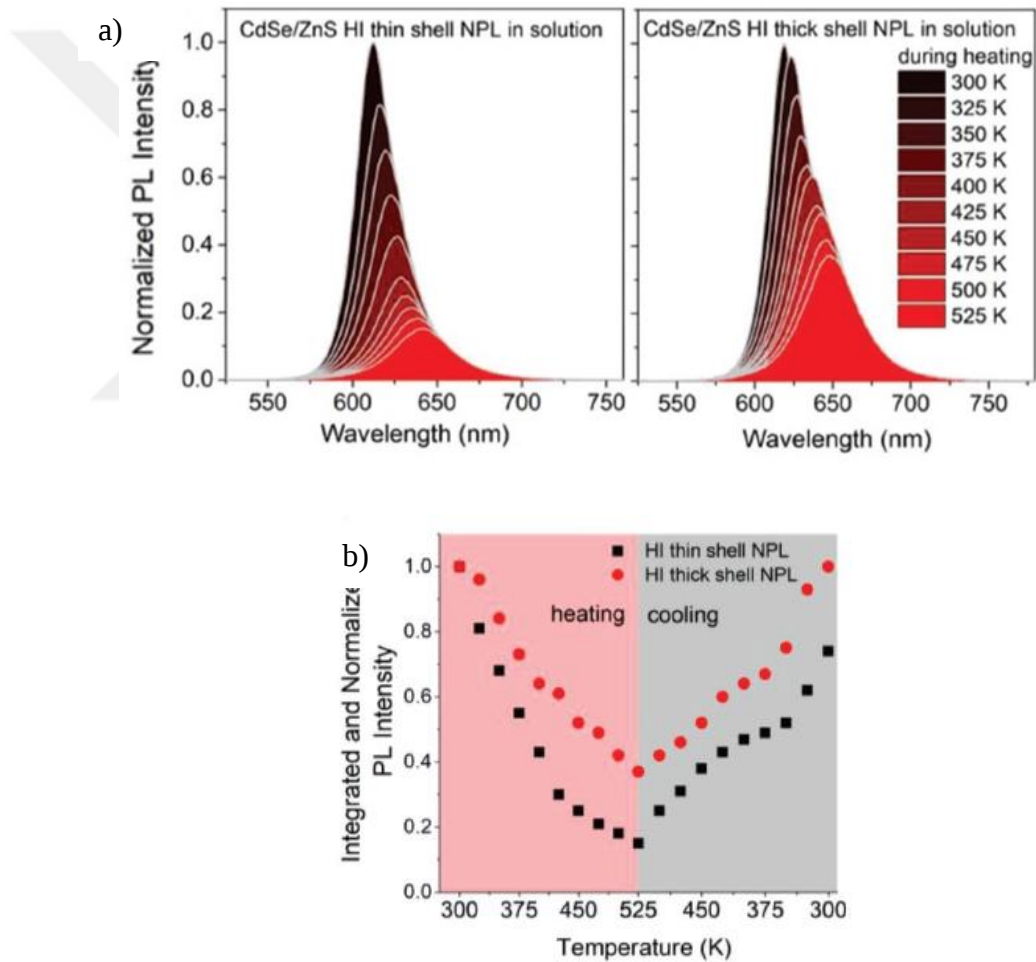


Figure 5.16: a) Normalized PL intensity values of colloidal CdSe/ZnS core/HI shell NPLs, both for the thin-shell and thick-shell samples, during temperature elevation from room temperature to 525 K under inert Ar gas and b) integrated PL intensity of the thin-shell and thick-shell CdSe/ZnS HI NPLs. Reprinted with permission from [2].

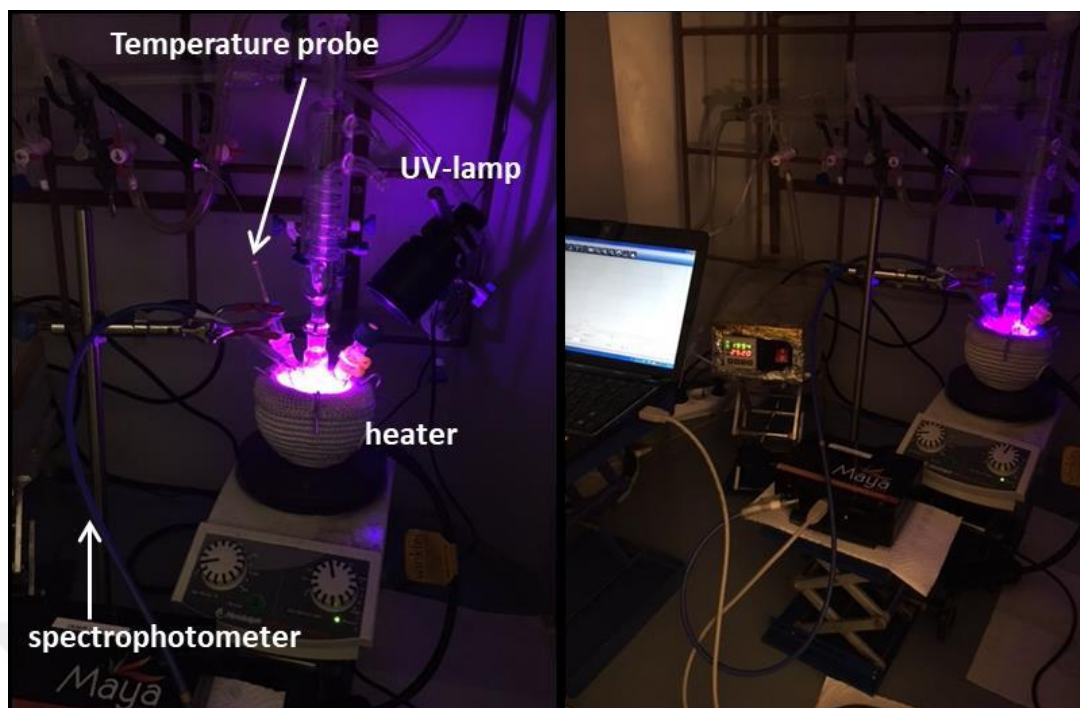


Figure 5.17: Images from our setup constructed to implement thermal tests in solution under inert gas.

Although colloidal tests provide important piece of information, semiconductor NCs are more commonly used in the form of thin films on substrates such as quartz. Therefore, the next meaningful step is to study the aging behaviour of NPLs deposited onto substrates by a spin-coater. The setup we constructed for solid film testing was similar to the previous one, with the exception of precise temperature control via the use of a thermal camera. The camera used is a FLIR-type thermal camera, which has $0.1\text{ }^{\circ}\text{C}$ precision. From Figure 5.18 it is apparent that the thick-shell sample is the most heat durable. All the samples exhibited red-shifts with temperature elevation. The c-ALD samples suffered a substantial increase in FWHM values, while for the HI shell heteronanoplatelets, this increase was small. The starting PL intensity of CdSe and CdSe/ZnS core/c-ALD shell drastically diminished when heating from 300 to 400 K. The exact PL intensity was 4%, 2% and 10% of the initial PL intensity for the core, thin-shell c-ALD and thick-shell c-ALD samples, respectively. On the other hand, the HI thin-shell and thick-shell NPLs maintained 40% and 52% of their PL intensities. At 450 K, the PL intensity of CdSe and CdSe/ZnS core/c-ALD shell samples was no longer detectable, whereas, the HI thin-shell and thick-shell samples preserved 20% and 35% of their respective PL emission. They even had 2% and 20% of their initial PL intensity at 500 K. Among all samples, our CdSe/ZnS HI thick-shell NPLs

demonstrated unparalleled thermal stability performance by fully recovering PL intensity during 300 to 450 K cycle in air.

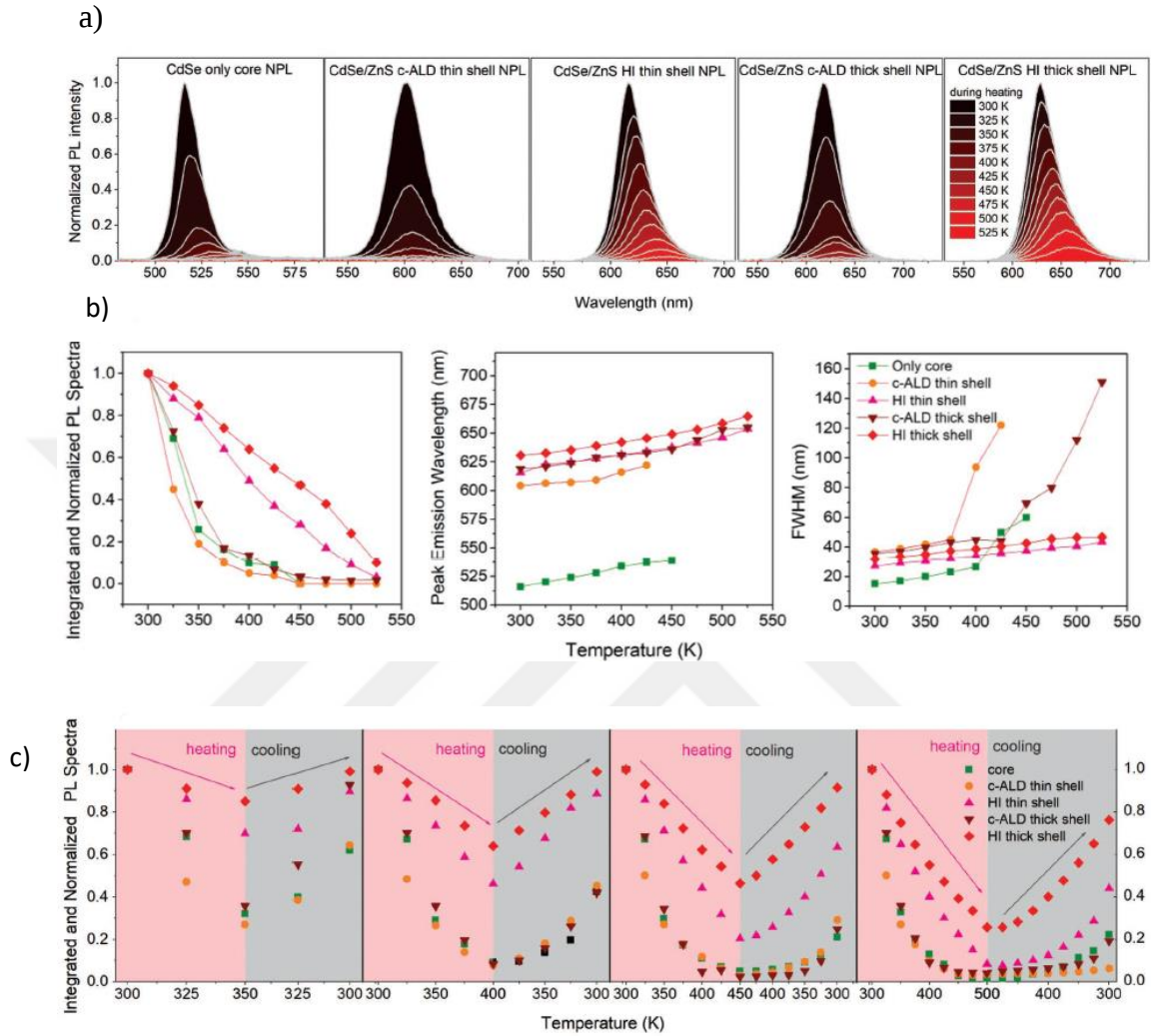


Figure 5.18: a) Normalized PL intensity values and b) the change of both normalized and integrated PL spectra, wavelength (nm) and FWHM values of deposited CdSe NPL, CdSe/ZnS HI thin-shell, CdSe/ZnS HI thick-shell, CdSe/ZnS c-ALD thin-shell and CdSe/ZnS c-ALD thick-shell samples while heating the quartz films from 300 to 525 K under ambient conditions, c) heating and cooling cycling of the same samples in the following temperature ranges: from 300 to 350 K; from 300 to 400 K; from 300 to 450 K and from 300 to 500 K. Reprinted with permission from [15].

5.3.2. Optical aging (UV test)

Another way of testing the endurance of the synthesized NPLs is to use UV exposure for definite intervals of time. It is known that both exposure to air and UV light causes

oxidation of semiconductor NCs [34],[35]. Usually, after long exposure to air and few-hour exposure to UV light, typically change in color, a red-shift in PL spectra, decrease in QY and decrease in PL intensity are observed. The change in both absorption and PL spectra happen due to the change (decrease) in the size of NCs coming from the loss of surface atoms and formation of new defect structures. It is known that oxygen molecules cause oxidation of chalcogenide atoms such as Se and S. This leads to the formation of SeO_2 and SO_4^{2-} [36]. These atoms may leave the surface of NCs, which leads to “dangling” Cd atoms in the case of CdSe QDs. The longer the exposure to UV light is, the more decomposition of NCs takes place. Therefore, performing UV illumination test unravels photostability properties of the tested samples.

In the case of our CdSe/ZnS core/HI shell NPLs, a UV-lamp with 354 nm peak emission was employed. The whole experiment lasted 48-hour, starting with smaller time intervals and continuing with longer ones (Figure 5.19). At the end of 48-hour non-interrupted exposure to UV light, although the thin-shell samples lost PL intensity completely, the thick-shell samples preserved 40% of their initial PL emission.

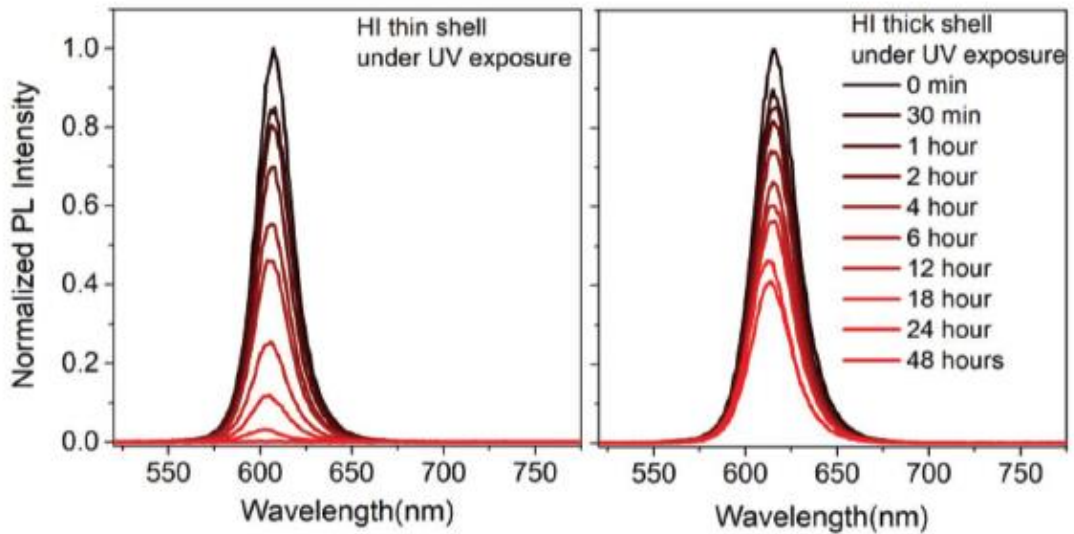


Figure 5.19: UV lamp illumination test results of CdSe/ZnS core/ shell NPL samples that were synthesized with our HI-shell method. Reprinted with permission from [2].

Chapter 6

Conclusion

Colloidal quantum wells exhibit unique features owing to their thickness-tunable optical properties and well-controlled atomically-flat surface areas. This family of nanocrystals displays step-like absorption profiles accompanied with narrow emission linewidths compared to QDs with giant oscillator strengths leading to high recombination rates. In this thesis, two main research efforts related to NPLs were made: First, to produce thermally and chemically durable NPLs, CdSe/ZnS heterostructures were synthesized in core/shell architecture, for the first time, employing hot-injection route. Second, to improve chemical stability and expand application possibilities of CdS NPLs, CdS were synthesized, which were unfolded using the combination of crown and c-ALD shell layers, while the QY of their core/crown heterostructures was also enhanced.

In CdSe/ZnS hot-injection shell growth, we proposed and synthesized CdSe/ZnS core/shell NPLs at 300°C under Ar gas using octanethiol-1 as the S precursor and Zn acetate as the Zn source. TEM images show that, when uniformly-shaped CdSe core NPLs were employed, the resultant CdSe/ZnS NPLs also possessed well-defined square shapes. Their QY was almost 100% in the case of thick-shell HI grown NPLs and it was sustained at 93% after 5-6 times of centrifugation with ethanol precipitating agent. Thermal and optical aging tests were carried out to analyze thermal and chemical endurance of the as-synthesized CdSe/ZnS hot-injection NPLs. In colloidal thermal test, at 400 K the thin-shelled NPLs maintained 42% of their initial PL intensity, while the thick-shelled ones preserved 65% of their initial PL intensity values. Also, at 525 K, 15% of the initial PL intensity of the thin-shelled was preserved, whereas 37% of the initial PL of the thick-shelled was conserved. In the case of spin-coated films at 450 K, the PL intensity of control group CdSe and CdSe/ZnS core/c-ALD shell samples was no longer detectable, whereas the HI thin-shell and thick-shell samples preserved 20% and 35% of their PL emission, respectively. The HI samples even had 2% and 20% of their initial PL intensity at 500 K. Among all samples,

CdSe/ZnS HI thick-shell NPLs demonstrated unparalleled thermal stability performance by fully recovering their PL intensity during a thermal cycle of 300 to 450 K in air. Therefore, these proposed NPLs hold great promise to find use in potential applications at elevated temperatures. After two days of exposure to a UV light source at 354 nm wavelength, the thick-shelled NPLs preserved 40% of their initial PL emission.

In CdS NPL synthesis, the trap emission was comparatively reduced to a negligible amount in the samples synthesized. For CdS/CdSe core/crown NPLs synthesized here, the results are promising in terms of the increased QY, almost non-existent trap emission, and enhanced chemical stability, which allows the cleaning process to be implemented by using a stronger polar solvent - ethanol. Both crown and c-ALD shell growths resulted in unfolding of CdS NPLs which, can be used for further enhancements if desired.

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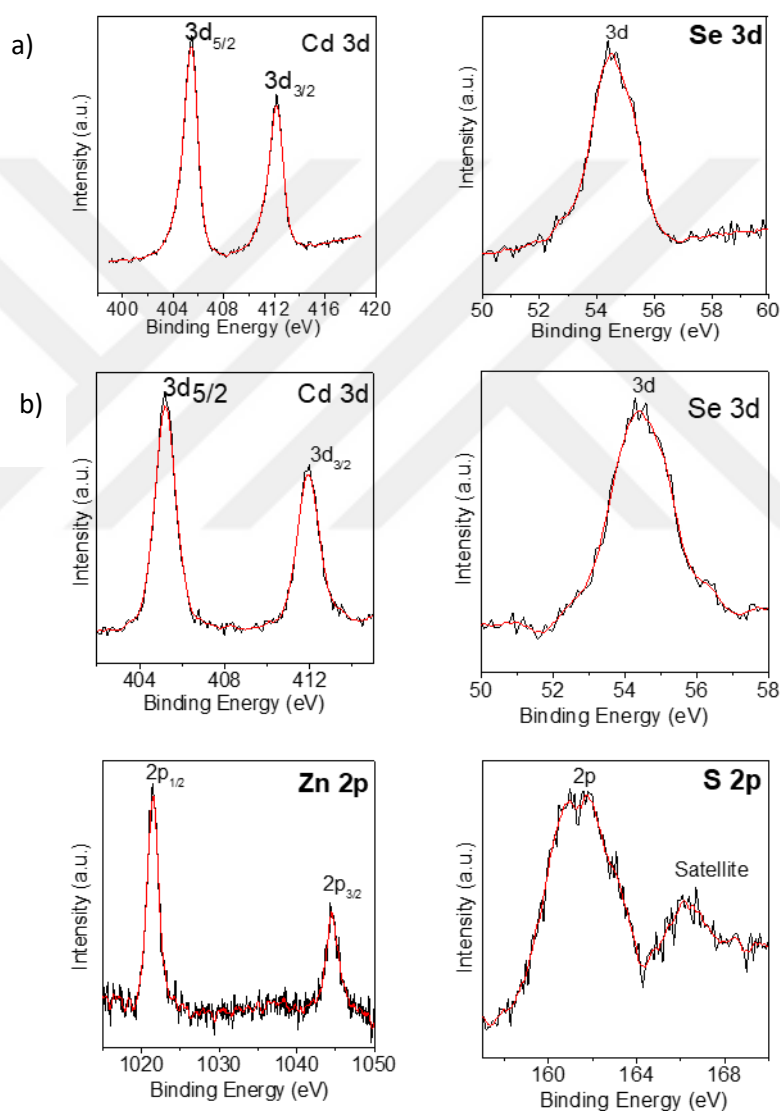
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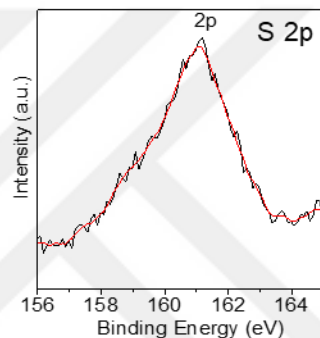
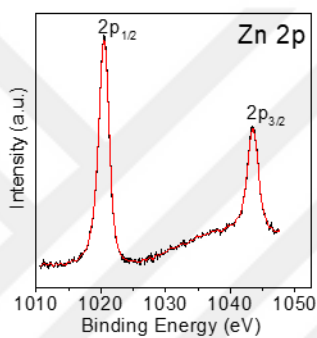
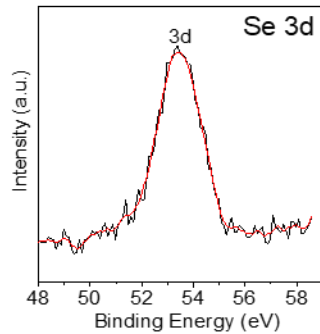
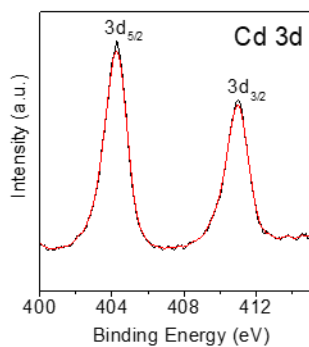


Appendix A

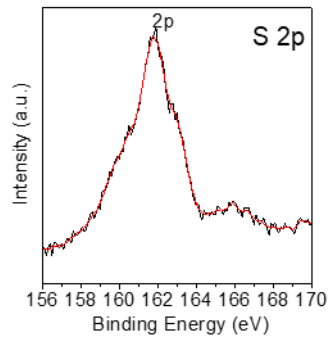
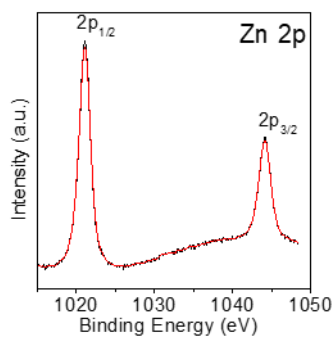
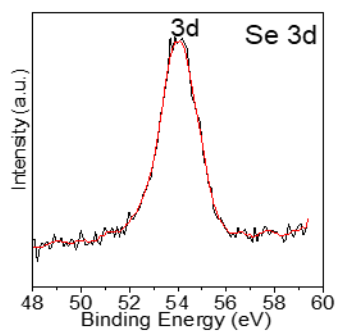
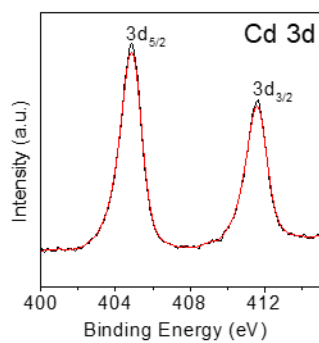
A.1. XPS spectroscopy



c)



d)



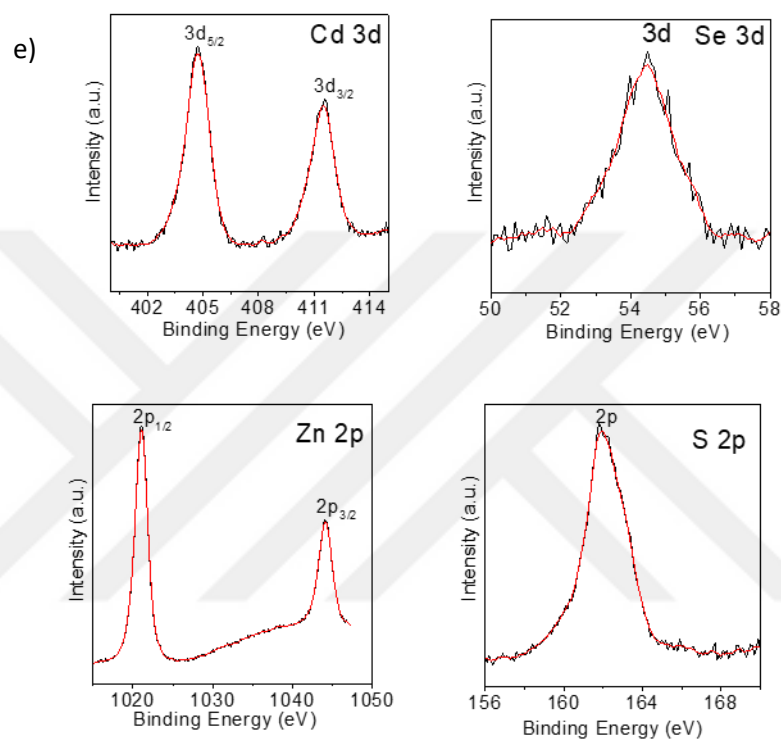


Figure A1: XPS data of the nanoplatelets synthesized in this thesis: a) CdSe 4 ML core NPLs, b) CdSe/ZnS thin c-ALD NPLs, c) CdSe/ZnS thick c-ALD NPLs, d) CdSe/ZnS hot-injection thin-shell NPLs, and e) CdSe/ZnS hot-injection thick-shell NPLs.